

## CHEM 1201 - Mastering Chemistry HW 1 & 2 + Exam 1 Review (Ch. 1-3, Prof. Elgrishi) study guide verified A+ tips NEW!!!

- 1. 2, 1:** What is the ratio of hydrogen atoms (HH) to oxygen atoms (OO) in 2 LL of water? Enter the simplest whole number ratio in order of hydrogen to oxygen, respectively.
- 2. 600 H atoms:** How many atoms of hydrogen (HH) are present in 200 molecules of ammonia ( $\text{NH}_3\text{NH}_3$ )?
- 3.  $8.35 \times 10^{25}$  H atoms:** A solution of ammonia and water contains  $2.90 \times 10^{25}$  water molecules and  $8.50 \times 10^{24}$  ammonia molecules. How many total hydrogen atoms are in this solution?
- 4. Intensive:**
  - matte grey in color
  - begins to melt at 961.78 degrees Celcius
  - has a density of 10.49 g/mL**Extensive:**
  - has a volume of 7.42 mL
  - has a mass of 77.83 g**Classify the following as intensive or extensive properties of silver.**
  - matte grey in color
  - begins to melt at 961.78 degrees Celcius
  - has a density of 10.49 g/mL
  - has a volume of 7.42 mL
  - has a mass of 77.83 g
- 5. the density of the sample**  
**the melting point of the sample:** You are given a sample resembling silver. What properties could be used to help determine whether the sample is really silver?
- 6. 189 g:** By definition,  
density=mass/volume  
Typical units of density are g/mL or  $\text{g/cm}^3$ .  
 $1\text{mL}=1\text{cm}^3$   
The density of iron is 7.87 g/mL. What is the mass of a  $4.00\text{ cm} \times 3.00\text{ cm} \times 2.00\text{ cm}$  block of iron?
- 7. In moist air it combines with oxygen to form aluminum oxide.**

# Mastering Chemistry Answers Ch 13

**Keith A. Sverdrup, Alyn C.  
Duxbury, Alison Duxbury**



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