

12. Molecular Geometry and Bonding Theories

Class: Principles of Chemistry, Semester II, 2019

Time: You should not spend the less than 45 minutes to complete this questionnaire (Total: 100)

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Please answer the Chapter 8 topics in **question 1** and **question 2** before proceeding for the other given assignments.

Original Outline	
<p>Learning Object: To understand hybridization and how it gives rise to different molecular geometries that various elements form. It is represented in Fig. 7 as conceptual diagrams that the various elements form different kinds of hybrid orbitals that depend on the number of atomic orbitals that combine. It has a hybrid orbital and the three unhybridized orbitals as p-orbitals. In addition, various orbitals form from combinations of orbitals in an overlap manner. Therefore, the hybrid orbitals form a new bond system. The resulting orbitals are called sp^3.</p> <p>Hybridization of the form:</p>	<p>Hybridization Diagram</p>
<p>Explain that are the combinations of atomic orbitals, a process.</p>	
<p>Question 1:</p> <p>What type of atomic orbitals are used in hybridization?</p>	
<p>Answer 1: Specify the electron configuration of atoms.</p>	<p>What are hybridized?</p>
<p>Answer: sp^3 hybrid orbitals</p>	
<p>Question 2: Further illustrate sp^3 hybridization theory for complex forms. Therefore, this will be relevant with the following question. Are hybrid orbitals and unhybridized orbitals are identical? Explain how orbitals form the type of hybridization?</p>	
<p>Answer 2: How to approach the problem?</p>	<p>What are hybridized?</p>

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H Kauffman



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