

Atomic Structure Worksheet

Name: Livingston Mengya

Period: 03

1. Name the three particles of the atom and their respective charges are:
 - a. Proton It has positive charge
 - b. Neutron It has neutral charge
 - c. Electron It has negative charge
2. The number of protons in one atom of an element determines the atom's Atomic number, and the number of electrons determines The number of shells of an element of an element.
3. The atomic number tells you the number of Protons in one atom of an element. It also tells you the number of electrons in a neutral atom of that element. The atomic number gives the "identity" of an element as well as its location on the Periodic Table. No two different elements will have the same atomic number.
4. The Atomic mass of an element is the average mass of an element's naturally occurring atoms, or isotopes, taking into account the Abundance of each isotope.
5. The Atomic mass of an element is the total number of protons and neutrons in the nucleus of the atom.
6. The mass number is used to calculate the number of neutrons in one atom of an element. In order to calculate the number of neutrons you must subtract the protons from the atomic mass.
7. Give the symbol and number of protons in one atom of:

Lithium <u>Li 3</u>	Bromine <u>Br 35</u>
Iron <u>Fe 26</u>	Copper <u>Cu 29</u>
Oxygen <u>O 8</u>	Mercury <u>Hg 80</u>
Arsenic <u>As 33</u>	Helium <u>He 2</u>

Matching Worksheet For Atomic Structure Answer Key

G Thomas



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 $\text{SO}_4^{2-} \text{ s} + 2 \text{ Ag}^+ \text{ aq} \rightleftharpoons 2 \text{ Ag}^+ \text{ aq} + \text{SO}_4^{2-} \text{ aq}$
 $\text{Al}(\text{OH})_3 \text{ s} + 3 \text{ Ag}^+ \text{ aq} \rightleftharpoons 3 \text{ Ag}^+ \text{ aq} + \text{Al}(\text{OH})_3 \text{ s}$
 $\text{Pb}(\text{OH})_2 \text{ s} + 2 \text{ Ag}^+ \text{ aq} \rightleftharpoons 2 \text{ Ag}^+ \text{ aq} + \text{Pb}(\text{OH})_2 \text{ s}$

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