# **Chapter 6**

# Chapter 6: Thermochemistry

# 6.2 - Enthalpy and Calorimetry

#### Enthalpy

$$H = E + PV$$

For a chemical reaction, enthalpy change is given by the equation

$$\Delta H = H_{products} - H_{reactants}$$

#### Products of a reaction have a greater enthalpy ( $\Delta H$ ):

- ΔH is positive
- · heat is absorbed by the system (endothermic reaction)

#### Reactants of a reaction have a greater enthalpy ( $\Delta H$ ):

- · AH is negative
- · heat is released by the system (exothermic reaction)

#### Calorimetry

#### Constant-pressure calorimetry

change in enthalpy = change in heat

$$\Delta H = q_p$$

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