CHEM 1201 - Mastering Chemistry HW 1 & 2 + Exam 1 Review (Ch. 1-3, Prof. Elgrishi) study guide verified A+ tips NEW!!!

- 1. 2, 1: What is the ratio of hydrogen atoms (HH) to oxygen atoms (OO) in 2 LL of water? Enter the simplest whole number ratio in order of hydrogen to oxygen, respectively.
- 2. 600 H atoms: How many atoms of hydrogen (HH) are present in 200 molecules of ammonia (NH3NH3)?
- 3. 8.35×10^25 H atoms: A solution of ammonia and water contains 2.90×1025 water molecules and 8.50×10^24 ammonia molecules. How many total hydrogen atoms are in this solution?
- 4. Intensive:
- matte grey in color
- begins to melt at 961.78 degrees Celcius
- has a density of 10.49 g/mL

Extensive:

- has a volume of 7.42 mL
- has a mass of 77.83 g: Classify the following as intensive or extensive properties of silver.

matte grey in color

begins to melt at 961.78 degrees Celcius

has a density of 10.49 g/mL

has a volume of 7.42 mL

has a mass of 77.83 g

5. the density of the sample

the melting point of the sample: You are given a sample resembling silver. What properties could be used to help determine whether the sample is really silver?

189 g: By definition,

density=mass/volume

Typical units of density are g/mL or g/cm³.

1mL=1cm^3

The density of iron is 7.87 g/mL. What is the mass of a 4.00 cm × 3.00 cm × cm × 2.00 cm block of iron?

7. In moist air it combines with oxygen to from aluminum oxide.

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