

Annotated Solution

2018 USNCO Local Exam

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Created on November 1st, 2020
Revised on April 7th, 2021

1 Solutions

1. We can use Avogadro's constant, and then do some dimensional analysis.

$$225 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32.0 \text{ g O}_2} \times \frac{2 \text{ mol O}}{1 \text{ mol O}_2} = \frac{6.92 \times 10^{23} \text{ atoms}}{1 \text{ mol O}} = \boxed{6.42 \times 10^{24} \text{ atoms}}$$

Thus, the answer is **C**.

2. Because the same number of moles must remain before and after dilution, we can set up the following equation:

$$\begin{aligned} M_1 V_1 &= M_2 V_2 \\ (1.50 \text{ M})(V_1) &= (0.200 \text{ M})(0.500 \text{ L}) \\ (1.50 \text{ M})(V_1) &= 0.100 \text{ M} \cdot \text{L} \\ V_1 &= 0.100 \text{ L} \end{aligned}$$

Thus is **100 mL**, or **A**.

3. We first need to write the reaction in order to perform the stoichiometry. The reaction is as follows:



Now we can do the stoichiometry:

$$10.00 \text{ g Cu} \times \frac{1 \text{ mol Cu}}{63.55 \text{ g Cu}} \times \frac{1 \text{ mol H}_2\text{O}}{2 \text{ mol Cu}} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{1.417 \text{ g H}_2\text{O}}$$

Thus, the answer is **B**.

4. Set a standard amount of sample at an arbitrary 100.0 g. Thus, there are 48.41 g Cu, 19.02 g As, and 32.57 g S. We can divide by the molar masses of each of the compounds:

$$\begin{aligned} 48.41 \text{ g Cu} &= \frac{1 \text{ mol Cu}}{63.55 \text{ g Cu}} = 0.7618 \text{ mol} \\ 19.02 \text{ g As} &= \frac{1 \text{ mol As}}{74.92 \text{ g As}} = 0.2539 \text{ mol} \\ 32.57 \text{ g S} &= \frac{1 \text{ mol S}}{32.06 \text{ g S}} = 1.016 \text{ mol} \end{aligned}$$

Dividing by the lowest number, there there will be the ratio of Cu : As : S = 3 : 1 : 4. Thus, the formula is **Cu₃AsS₄**, or **C**.

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