Question 30

A student performs a titration with a base as the titrant and an acid as the analyte. He finds that the pH at the equivalence point is 8.3, and concludes that he was adding NaOH to HF. Which of the following observations made during the titration does NOT support his conclusion?

A. The $[H^+]$ concentration in the experimental vessel was higher before the titration began than at the equivalence point

You did not select. This answer is incorrect

B. The solution, at the equivalence point, was found to be a conductor of electricity

You did not select. This answer is incorrect

C. To reach the equivalence point, 10 mL of base (0.5 M) was added to 5 mL of acid (0.25 M)

You did not select. This answer is correct

D. When the student titrates past the equivalence point, the solution reaches a pOH of 1

This answer is incorrect

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