

Half Modern Chemistry Review

CHAPTER 9: STOICHIOMETRY

The following pages contain the book (but not all) of the information for the chapter 9 test.

Focus on this content, but make sure to review class notes, activities, assessments, questions, etc.

If you study this document and NOTHING else, you should at least be able to PASS the test.

**** Test items will be recall, application, and/or application of this content. ****

OUTCOMES

- Calculate with percent in stoichiometry context. (C, C)
- Calculate stoichiometry context in reaction and percent (B, C)
- 9.1. Determine moles of reactants from balanced chemical equations. (T & P)
- 9.2. Perform stoichiometry calculations with: a) mole to mole, mole to gram, gram to mole, unknown to gram. (P, L, P)
- 9.3. Calculate stoichiometry problems and determine the limiting reagent, excess reagents, and amount of product produced. (P & P)

9.1: INTRODUCTION TO STOICHIOMETRY

a. Introductory

- stoichiometry: stoichiometry -- calculations involving the mass relationships of elements in compounds.
- reaction stoichiometry -- calculations involving the mass relationships between reactants and products in a chemical reaction.
- molar ratio -- a conversion factor that relates the amount, in moles, of any two substances involved in a chemical reaction.

b. Chapter Highlights

- Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction.
- Stoichiometry is based on the law of conservation of mass.
- Relating one substance to another requires recognizing the amount of each substance in moles.
- A molar ratio is the conversion factor that relates the amount in moles of any two substances in a chemical reaction. The molar ratio is derived from the balanced equation.
 - Example: $2\text{FeCl}_3 + 3\text{Cl}_2 \rightarrow 2\text{FeCl}_4 + 3\text{Cl}_2$ What is the ratio of iron to chlorine moles? a) 2 to 3
- Amount of a substance is measured in moles, and mass of a substance is determined using molar mass such as grams, kilograms, or milligrams.
- Mass and amount of substances are quantities, substances moles and grams are units.
- A balanced chemical equation is necessary to solve any stoichiometric problem.

9.2: BASIC STOICHIOMETRIC CALCULATIONS

a. Chapter Highlights

- In an ideal stoichiometric calculation, the mass, or the amount of any reactant or product can be calculated if the balanced chemical equation and the mass or amount of any other reactant or product is known.

9.3: LIMITING REACTANTS AND PERCENTAGE YIELD

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Katrin Zwirgmaier



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