

# Unit Review

1. Calculate the molar mass of organic compounds:

a)  $\text{HCl}$

b)  $\text{H}_2\text{SO}_4$

c)  $\text{FeCl}_2\text{O}_4$

a)  $\text{H}_2\text{O}$

b)  $\text{Cu}_2\text{OH}_2$

d)  $\text{Fe}_2\text{SO}_4$

a)  $\text{H}_2\text{SO}_4$

b)  $\text{Fe}_2\text{SO}_4$

c)  $\text{CH}_3\text{COOH}$

a)  $\text{CO}_2$

b)  $\text{FeCl}_2$

d)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$

a)  $\text{CH}_4$

b)  $\text{FeCl}_3$

2. Calculate the mass of the following:

a) 1.00 moles of  $\text{H}_2\text{SO}_4$

c) 0.00500 moles of  $\text{Fe}_2\text{SO}_4$

b) 0.001 moles of  $\text{H}_2\text{SO}_4$

d) 0.001 moles of  $\text{FeCl}_3$

a) 0.001 moles of  $\text{FeCl}_3$

3. Calculate the number of moles in the following:

a) 17.0 g of  $\text{H}_2\text{SO}_4$

c) 0.100 g of  $\text{CO}_2$

b) 17.0 g of  $\text{H}_2\text{O}$

d) 0.100 g of  $\text{CH}_4$

a) 0.001 g of  $\text{C}$

4. Calculate the molar mass of organic compounds and then calculate the mass of the following:

a) A 0.001 mol sample of  $\text{H}_2\text{SO}_4$  has a mass of 0.001 g

b) A 0.00500 mol sample of  $\text{Fe}_2\text{SO}_4$  has a mass of 0.00500 g

c) The mass of 0.001 mol of  $\text{FeCl}_3$  is 0.001 g

d) A 0.001 mol sample of  $\text{FeCl}_3$  has a mass of 0.001 g

5. Calculate the number of moles of the following:

a) 17.0 g of  $\text{H}_2\text{SO}_4$

c) 0.100 g of  $\text{CO}_2$

b) 0.100 g of  $\text{CH}_4$

d) 0.100 mol of  $\text{CO}_2$

a) 0.001 mol of  $\text{FeCl}_3$

c) 0.100 mol of  $\text{CO}_2$

6. Calculate the volume of the following gases at STP:

a) 0.001 mol of  $\text{H}_2$

c) 0.001 mol of  $\text{CO}_2$

# Mole Review Answers

**Morris Hein, Susan Arena**



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Foundations of College Chemistry, Alternate Morris Hein, Susan Arena, 2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis      *Reviews of College Board Tests from the Fifth Mental Measurements Yearbook* ,1959      **Foundations of College Chemistry** Morris Hein, Susan Arena, Cary Willard, 2023 Foundations of College Chemistry 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence The central focus is to make chemistry interesting and understandable and teach students the problem solving skills they will need This International Adaptation offers new and updated content with improved presentation of all course material It builds on the strengths of previous editions including clear explanations and step by step problem solving The material emphasizes real world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry There is new and expanded coverage of polarizing power and polarizability Fajans rules collision number and mean free path abnormal molecular masses and van t Hoff factor and applications of radioactivity      Foundations of College Chemistry Morris Hein, Susan Arena, 2013-01-01 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years Foundations of College Chemistry Alternate 14th Edition has helped readers master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis      **The Inclusive Classroom** Margo A. Mastropieri, Thomas E. Scruggs, 2010 This text offers a wealth of practical and proven strategies for successfully including students with disabilities in general education classrooms The text provides targeted strategies for the subject and skill areas as well as special needs of individual students with a strong focus on instructional strategies applied to specific student need areas An overall theme of effective differentiated instruction is infused throughout the text relating to those practices that are most closely aligned with academic success The text is unique in its three part coverage of 1 the fundamentals of teaching students with special needs 2 effective general teaching practices and 3 inclusive practices in specific subject areas With a strong focus on instructional strategies and how they are applied to specific student need areas the text goes further by featuring more specific strategies

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*Basic Concepts of Chemistry, Study Guide* Leo J. Malone,1989-02-21 This Third Edition revised to provide smoother transitions between topics employs a concise yet informal approach to basic chemistry organized to help students employ basic math skills and problem solving strategies Writing style is straightforward and presentation incorporates many concrete analogies to clarify new concepts Includes many illustrative worked examples

**Student Solutions Manual** Darrell Ebbing,2005-03-17 Provides solutions to odd numbered Practice Problems General Problems and Cumulative Skills Problems plus answers to Review Questions Luther League Review ,1928 *The American Review of Reviews* Albert Shaw,1923 **American Monthly Review of Reviews** Albert Shaw,1928 **Hygeia** ,1923 **The Poetry Review** ,1987

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