

Chemistry 12 – Unit I – Reaction Kinetics  
Kinetics Review Test

1. Consider the following reaction mechanism:

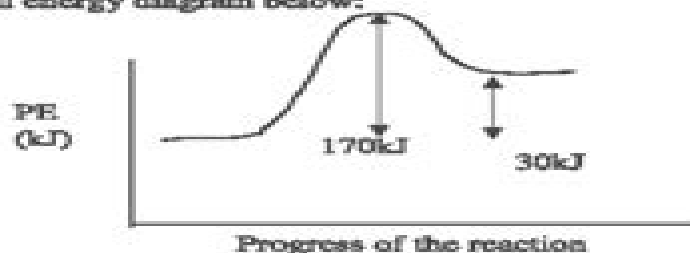


The chemical species MX is a(n) \_\_\_\_\_.

- A. catalyst                      B. inhibitor                      C. product                      D. reaction intermediate
2. Consider the following reaction:  

$$2 \text{N}_2\text{O}_5(g) \rightarrow 4 \text{NO}_2(g) + \text{O}_2(g)$$
 At a certain temperature the rate of decomposition of  $\text{N}_2\text{O}_{5(g)}$  is  $2.5 \times 10^{-6} \text{ mol/s}$ . The rate of formation of  $\text{NO}_2$  is \_\_\_\_\_.
- A.  $1.0 \times 10^{-5} \text{ mol/s}$     B.  $1.3 \times 10^{-6} \text{ mol/s}$     C.  $2.5 \times 10^{-6} \text{ mol/s}$     D.  $5.0 \times 10^{-6} \text{ mol/s}$
3. Which of the following factors affect the rates of *both* homogeneous and heterogeneous reactions?
- |                                |                                   |
|--------------------------------|-----------------------------------|
| I      nature of the reactants | II      presence of a catalyst    |
| III     temperature of system  | IV     concentration of reactants |
- A. I and IV only              B. II and III only              C. II, III, and IV only              D. I, II, III, and IV
4. Which of the following equations represents an endothermic reaction?
- A.  $\text{N}_2\text{O}_{5(g)} + 59 \text{ kJ} \rightarrow 2 \text{NO}_{2(g)}$               B.  $2 \text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2 \text{H}_2\text{O}_{(g)} + 572 \text{ kJ}$   
 C.  $2 \text{BrCl}_{(g)} - 29.3 \text{ kJ} \rightarrow \text{Br}_{2(g)} + \text{Cl}_{2(g)}$               D.  $2 \text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2 \text{H}_2\text{O}_{(g)} ; \Delta H = -572 \text{ kJ}$

5. Consider the potential energy diagram below:



The activation energy for the reverse reaction is \_\_\_\_\_.

- A. 30 kJ                      B. 140 kJ                      C. 170 kJ                      D. 200 kJ

# Modern Chemistry Ch 17 Reaction Kinetics Test

**G.H. Duffey**



## **Modern Chemistry Ch 17 Reaction Kinetics Test:**

*Modern Physical Chemistry* G.H. Duffey, 2013-11-11 In this new textbook on physical chemistry fundamentals are introduced simply yet in more depth than is common Topics are arranged in a progressive pattern with simpler theory early and more complicated theory later General principles are induced from key experimental results Some mathematical background is supplied where it would be helpful Each chapter includes worked out examples and numerous references Extensive problems review and discussion questions are included for each chapter More detail than is common is devoted to the nature of work and heat and how they differ Introductory Caratheodory theory and the standard integrating factor for dGrev are carefully developed The fundamental role played by uncertainty and symmetry in quantum mechanics is emphasized In chemical kinetics various methods for determined rate laws are presented The key mechanisms are detailed Considerable statistical mechanics and reaction rate theory are then surveyed Professor Duffey has given us a most readable easily followed text in physical chemistry Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 **Modern Chemistry** Raymond E. Davis, 1999 2000 2005 State Textbook Adoption Rowan Salisbury

**Physics Briefs**, 1982 *Publications of the National Institute of Standards and Technology ... Catalog* National Institute of Standards and Technology (U.S.), National Institute of Standards and Technology (U.S.). Information Resources and Services Division, 1994 Modern Electrochemistry 2A John O'M. Bockris, Amulya K.N. Reddy, Maria E. Gamboa-Aldeco, 2001-01-31 This book had its nucleus in some lectures given by one of us J O M B in a course on electrochemistry to students of energy conversion at the University of Pennsylvania It was there that he met a number of people trained in chemistry physics biology metallurgy and materials science all of whom wanted to know something about electrochemistry The concept of writing a book about electrochemistry which could be understood by people with very varied backgrounds was thereby engendered The lectures were recorded and written up by Dr Klaus Muller as a 293 page manuscript At a later stage A K N R joined the effort it was decided to make a fresh start and to write a much more comprehensive text Of methods for direct energy conversion the electrochemical one is the most advanced and seems the most likely to become of considerable practical importance Thus conversion to electrochemically powered transportation systems appears to be an important step by means of which the difficulties of air pollution and the effects of an increasing concentration in the atmosphere of carbon dioxide may be met Cor sion is recognized as having an electrochemical basis The synthesis of nylon now contains an important electrochemical stage Some central biological mechanisms have been shown to take place by means of electrochemical reactions A number of American organizations have recently recommended greatly increased activity in training and research in electrochemistry at universities in the United States **Fossil Energy Update**, 1986 *Nuclear Science Abstracts*, 1975-04 Report summaries United States. Environmental Protection Agency, 1983

**Decennial Index to Chemical Abstracts**, 1926 **Modern Electrochemistry** John O'M. Bockris, 1998 **Analytical**

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