You are provided with the following:

- 3.3g metal carbonate, MCO<sub>3</sub>, labeled solution Q
- · 2M hydrochloric acid, labeled solution P
- Sodium hydroxide, labeled solution R containing 40g/L of solution

You are required to determine the relative atomic mass of metal M

#### Procedure

- Measure accurately 100cm<sup>3</sup> of solution P into clean 250cm<sup>3</sup> conical flak and add all the 3.3g of solid Q, MCO<sup>3</sup>
- ii. Shake the mixture well and wait for effervescence to stop. Label the resulting solution as S
- Pipette 25cm<sup>3</sup> of solution R into a conical flask and add 2-3 drops of phenolphthalein indicator.
- Fill the burette with solution S and titrate against the solution R until the end point.
- Record your results in the table below. Repeat the procedure at least two times to complete the table.
  (4 mks)

Annones and the same of the sa	i i	iii	Liii	
Final burette reading (cm <sup>3</sup> )				
Initial burette reading (cm3)				
Volume of solution S used (cm <sup>3</sup> )				

a) What is the average volume of solution S used?

(Timak)

- b) Calculate the moles of sodium hydroxide, solution R used. (2mks)
- c) Calculate the moles of hydrochloric acid in the average volume of solution S used. (2mks)

# June 2013 Chemistry Paper 3

**Jin-Ying Zhang** 

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