



Exploring Equilibrium

It Works Both Ways

Introduction.

The sound equilibrium has two rooms work meaning result, and Alma. receiving weight or balance. Our physical sense of equilibrium—in the motion of a sector or the swing of a pendulum-suggests as equal balance of apposing forces. How does this physical sense of equilibrium translate to chemical equilibrium? Explore the nature and consequences of equilibrium in chemical reactions.

Concepts

Review-thile magnitudes.

Chemical equilibrium

Commission reaction.

Activishment medications.

Background

Physical changes, such as nielting for or classifying sugar, are often increduced by noting that these processes can be easily reserved. Some consnone reamples of chemical charge, such as burning wood or spoiling. find, generally cannot be reserved. A closer look at chemical change, have even neverals that, many observical reactions are also reversible.

Consider the following evanuels of a revenible chemical reaction (Resignor 1). As both pressures and in the presence of a catalog, repregen and hydrogen read to form ammonia. If the temperature is high enough, however, ananosni a deconomoses to reform its constituent elements. The reaction can so both ward This reverelle reaction is represented symbolically using double arrenes.

Physical Science 2:

What happens when entrogen and budgen are allowed to reach In a allowed wosterns, the concentrations of nitrogens and hydrogen will decrease, and the concentration of anymonia will steadily increase as the traction processes in the forward direction. Soon, however, the concentration of annuous a will be large enough that the reverse reaction will begin to take place at a significant rate as well. Ecentrally, as the resction occurs in both the forecast and the reverse directions, the mander of anymorita modesiales being formed will become estad to the mander of animonia molecules being consumed. At this point, no further chargeex will be observed in the overall concentrations of nigrages, hydrogenand arrandonia. This is the point of chemical equilibrium.

Obenited equilibrium is defined as the state in which the rate of the for each reaction equals the rate of the receive reaction and the concerntrailores of resitants and tresducts remain constant, with time. Note that this definition describes a dynamic picture of equilibrium. The reactions continue, but there is an equal balance of opposing reaction rates.

What have no when the equal bean to disparted? Any Gater that changes the rate of the forward or the reverse reaction will change the amongings of egactanes and provided a that are preparet, at equilibrium. Reaction conditions that are known to affect the rates of chemical reactions trefude the concentrations of reactants and the temperature. In this experiment, we will investigate how changes in reaction conditions affect the amounts of reactants and products present at equilibrium.

Experiment Overview

The purpose of this experiment is to explore the nature of chemical equals became and to sciences, conditions that allect the exection of equifiltrians. Two different reversible reactions will be studied.

Reaction of iron(IIII) nitrate with potassium thiocyanics will be used to study constitute for equilibrium. Trans(IIII) tons read, with abhacyanate tons to firm FeSCN* complex ions (Equation 2). The effects of changing the concentrations of reactions and of changing the reaction terrementane will be investigated.

Pignation 2

The properties of an inclinator will be used to study and Astropolitical since. Any traditionation for a class others, care god to on lover by divergent forms are flower. substances that have different adors. Equation 3 supercentures the reversthis reaction of the incidenter become read green (HIn). Him represents an uncharged indicator molecule and In-antroductor anion formed after the neclecule has lost a hydrogen ion. The color of the indicator in the presence of either excess H. or OH. torus here Equation 4) will show how changing the concentration of a product affects the equilibrium absences for Economics 3.

Primary Class 3

$$H(ac) + OH(ac) \rightarrow H,O00$$

Paparathian 4

<u>Lab Report On Exploring Equilibrium It Works Both</u> <u>Ways</u>

Judd E. Hollander

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