



## Limiting Reactant Identification

6.7 moles of iron (Fe) react with 8.4 moles of oxygen ( $\text{O}_2$ ) according to the balanced equation above. What is the limiting reactant?

Iron (Fe) gmm:

$$(1)(55.9) = 55.9 \text{ g/mol}$$

Oxygen ( $\text{O}_2$ ) gmm:

$$(2)(16.0) = 32.0 \text{ g/mol}$$

mole ratio

$$\frac{6.7 \text{ mol Fe} \times 2 \text{ mol Fe}_2\text{O}_3}{4 \text{ mol Fe}} = 3.35 \text{ mol Fe}_2\text{O}_3$$

Limiting

Exces

$$\frac{8.4 \text{ mol O}_2 \times 2 \text{ mol Fe}_2\text{O}_3}{3 \text{ mol O}_2} = 5.6 \text{ mol Fe}_2\text{O}_3$$

# Limiting Reactant Problems And Answers High School

**DJ Losen**



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