

Preparation and Standardization of a NaOH Solution and Determination of the Molar Mass of an Unknown Acid Lab Report

This experiment involves two separate acid-base titration procedures. The purpose of the first one, Preparation and Standardization of a NaOH Solution is to perform an acid-base titration, in order to determine the equivalence point of acid and base. After the acid has been neutralized by the base, the results can be used to determine the concentration of sodium hydroxide solution by using stoichiometry. For the second experiment of Determination of the Molar Mass of an Unknown Acid it was to perform an acid-base titration, in order to find the molarity and molar mass of unknown acid. The second experiment uses the concentration of NaOH solution found from the first experiment.

Titration is an experiment used often for determining the concentration of an acid and base solution through the neutralization of the standard solution of base or acid. The major point in titration is to obtain the equivalence point. The equivalence point is when the acid and base have been neutralized. it is at the equivalence point where the number of moles of NaOH and $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}_{(s)}$ are reached. At the equivalence point, the base NaOH and the acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}_{(s)}$ have canceled each other out. This can be demonstrated by the chemical formula: $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}_{(s)} + 2 \text{NaOH}_{(aq)} \rightarrow \text{Na}_2\text{C}_2\text{O}_{4(aq)} + 2 \text{H}_2\text{O}_{(aq)}$ (Titration). When the acid and base are completely neutralized and the reaction is complete, the titration has reached the endpoint. The endpoint is indicated by an indicator, which is a substance that changes color to mark the endpoint. For this experiment Phenolphthalein is the indicator, it turns light pink to show that base and acid have been neutralized and is at endpoint. Standardization is the method of knowing the exact concentration of a solution. The standard solution is the base or acid that the concentration is known. Standardization was necessary in this experiment so that the concentration of base and acid are accurate. using just the weights of base or acid doesn't give an

Lab Report For An Unknown Salt Solution

Ralph A. Burns



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