

Name: _____

Date: _____

Student Exploration: Limiting Reactants

Vocabulary: chemical equation, chemical formula, chemical reaction, coefficient, limiting reactant, molecule, product, reactant, subscript

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

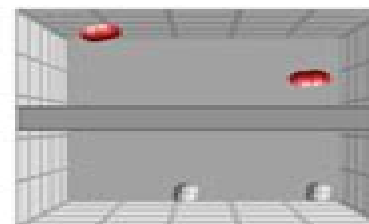
Imagine you and your friends are making hot dogs. A complete hot dog consists of a wiener and a bun. At the store, you buy four packages of eight wieners and three bags of 10 buns.

1. How many total hot dogs can you make?
_____ 30 _____
2. Which ingredient limited the number of hot dogs you could make?
_____ buns _____
3. Which ingredient will you have leftovers of?
_____ hotdogs _____

Gizmo Warm-up

Just as ingredients can be put together to make a new food, substances can combine during a **chemical reaction** to produce new substances. The substances that undergo change are called **reactants**. The new substances are **products**.

Sometimes during a chemical reaction, one type of reactant will be used up before the other reactants. This reactant is the **limiting reactant**. Using the *Limiting Reactants Gizmo™*, you can determine which reactant is limiting in various scenarios.



To begin, make sure $\text{H}_2 + \text{O}_2$ becomes H_2O is selected. The small "2" in H_2 , O_2 , and H_2O is a **subscript**. Subscripts represent the number of atoms in a **molecule**.

1. Use the sliders to set the number of O_2 molecules and H_2 molecules to two.
 - A. _____ How many hydrogen molecules (H_2) are there? _____ 2 _____ Hydrogen atoms? _____ 4 _____
 - B. _____ How many oxygen molecules (O_2) are there? _____ 2 _____ Oxygen atoms? _____ 4 _____



Limiting Reactant Gizmo Answers

Lifelique



Limiting Reactant Gizmo Answers:

Limiting Reactant and Percent Lifeliqe,2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

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Table of Contents Limiting Reactant Gizmo Answers

1. Understanding the eBook Limiting Reactant Gizmo Answers
 - The Rise of Digital Reading Limiting Reactant Gizmo Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactant Gizmo Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reactant Gizmo Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactant Gizmo Answers
 - Personalized Recommendations
 - Limiting Reactant Gizmo Answers User Reviews and Ratings
 - Limiting Reactant Gizmo Answers and Bestseller Lists

5. Accessing Limiting Reactant Gizmo Answers Free and Paid eBooks
 - Limiting Reactant Gizmo Answers Public Domain eBooks
 - Limiting Reactant Gizmo Answers eBook Subscription Services
 - Limiting Reactant Gizmo Answers Budget-Friendly Options
6. Navigating Limiting Reactant Gizmo Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reactant Gizmo Answers Compatibility with Devices
 - Limiting Reactant Gizmo Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reactant Gizmo Answers
 - Highlighting and Note-Taking Limiting Reactant Gizmo Answers
 - Interactive Elements Limiting Reactant Gizmo Answers
8. Staying Engaged with Limiting Reactant Gizmo Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reactant Gizmo Answers
9. Balancing eBooks and Physical Books Limiting Reactant Gizmo Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reactant Gizmo Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reactant Gizmo Answers
 - Setting Reading Goals Limiting Reactant Gizmo Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reactant Gizmo Answers
 - Fact-Checking eBook Content of Limiting Reactant Gizmo Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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