

## Limiting Reactant and Percent Yield Worksheet

(Show your work)

1. Consider the following reaction:  $2 \text{Al} + 6 \text{HBr} \rightarrow 2 \text{AlBr}_3 + 3 \text{H}_2$

When 3.22 moles of Al reacts with 6.96 moles of HBr, what are the limiting and excess reactants?

2. Consider the following reaction:  $4 \text{FeS}_2 + 11 \text{O}_2 \rightarrow 2 \text{Fe}_2\text{O}_3 + 8 \text{SO}_2$

When 26.62 moles of  $\text{FeS}_2$  reacts with 59.44 moles of  $\text{O}_2$ , what are the limiting and excess reactants?

3. Consider the following reaction:  $3 \text{Si} + 2 \text{N}_2 \rightarrow \text{Si}_3\text{N}_4$

When 600 g of Si reacts with 500 g of  $\text{N}_2$ , What are the limiting and excess reactants?

4. Given the following equation:  $\text{Al}_2(\text{SO}_4)_3 + 6 \text{NaOH} \rightarrow 3 \text{Na}_2\text{SO}_4 + 2 \text{Al}(\text{OH})_3$

If 10.0 g of  $\text{Al}_2(\text{SO}_4)_3$  is reacted with 10.0 g of NaOH, determine the limiting and excess reactants.

# Limiting Reactant And Percent Yield Review Sheet 3 Answers

**Kenneth W. Raymond**



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