

Limiting Reactant

For the following reactions, find the following:

- Which of the reactants is the limiting reactant?
 - What is the maximum amount of each product that can be formed?
 - How much of the excess reactant is left over after the reaction is complete?
- See next page for calculations!

1. Consider the following reaction:



a) limiting reactant

Answer the questions above, assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate.

$$\begin{array}{l}
 \frac{30.0 \text{ g NH}_4\text{NO}_3}{1} \times \frac{1 \text{ mol NH}_4\text{NO}_3}{80.0 \text{ g NH}_4\text{NO}_3} \times \frac{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4}{3 \text{ mol NH}_4\text{NO}_3} \times \frac{149.12 \text{ g } (\text{NH}_4)_3\text{PO}_4}{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4} = 18.6 \text{ g } (\text{NH}_4)_3\text{PO}_4 \\
 \frac{50.0 \text{ g Na}_3\text{PO}_4}{1} \times \frac{1 \text{ mol Na}_3\text{PO}_4}{162.94 \text{ g Na}_3\text{PO}_4} \times \frac{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4}{1 \text{ mol Na}_3\text{PO}_4} \times \frac{149.12 \text{ g } (\text{NH}_4)_3\text{PO}_4}{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4} = 45.5 \text{ g } (\text{NH}_4)_3\text{PO}_4 \\
 \frac{30.0 \text{ g NH}_4\text{NO}_3}{1} \times \frac{1 \text{ mol NH}_4\text{NO}_3}{80.0 \text{ g NH}_4\text{NO}_3} \times \frac{3 \text{ mol NaNO}_3}{3 \text{ mol NH}_4\text{NO}_3} \times \frac{85.00 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 31.9 \text{ g NaNO}_3
 \end{array}$$

2. Consider the following reaction:



a) limiting reactant

Answer the questions at the top of this sheet, assuming we start with 100.0 grams of calcium carbonate and 45.0 grams of iron (II) phosphate.

$$\begin{array}{l}
 \frac{100.0 \text{ g CaCO}_3}{1} \times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{3 \text{ mol CaCO}_3} \times \frac{310.18 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} = 103.3 \text{ g Ca}_3(\text{PO}_4)_2 \\
 \frac{45.0 \text{ g FePO}_4}{1} \times \frac{1 \text{ mol FePO}_4}{150.82 \text{ g FePO}_4} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{2 \text{ mol FePO}_4} \times \frac{310.18 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} = 46.3 \text{ g Ca}_3(\text{PO}_4)_2 \\
 \frac{45.0 \text{ g FePO}_4}{1} \times \frac{1 \text{ mol FePO}_4}{150.82 \text{ g FePO}_4} \times \frac{1 \text{ mol Fe}_2(\text{CO}_3)_3}{2 \text{ mol FePO}_4} \times \frac{291.73 \text{ g Fe}_2(\text{CO}_3)_3}{1 \text{ mol Fe}_2(\text{CO}_3)_3} = 43.5 \text{ g Fe}_2(\text{CO}_3)_3
 \end{array}$$

Limiting Reactant Review Sheet Key

Kenneth W. Raymond



Limiting Reactant Review Sheet Key:

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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Foundations of College Chemistry, Alternate Morris Hein, Susan Arena, 2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They'll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis

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General Chemistry Key Concise Review Guide Sixth Edition Darrell D. Ebbing, 1999

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Table of Contents Limiting Reactant Review Sheet Key

1. Understanding the eBook Limiting Reactant Review Sheet Key
 - The Rise of Digital Reading Limiting Reactant Review Sheet Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactant Review Sheet Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reactant Review Sheet Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactant Review Sheet Key
 - Personalized Recommendations
 - Limiting Reactant Review Sheet Key User Reviews and Ratings
 - Limiting Reactant Review Sheet Key and Bestseller Lists
5. Accessing Limiting Reactant Review Sheet Key Free and Paid eBooks
 - Limiting Reactant Review Sheet Key Public Domain eBooks
 - Limiting Reactant Review Sheet Key eBook Subscription Services
 - Limiting Reactant Review Sheet Key Budget-Friendly Options
6. Navigating Limiting Reactant Review Sheet Key eBook Formats

- ePub, PDF, MOBI, and More
 - Limiting Reactant Review Sheet Key Compatibility with Devices
 - Limiting Reactant Review Sheet Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reactant Review Sheet Key
 - Highlighting and Note-Taking Limiting Reactant Review Sheet Key
 - Interactive Elements Limiting Reactant Review Sheet Key
 8. Staying Engaged with Limiting Reactant Review Sheet Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reactant Review Sheet Key
 9. Balancing eBooks and Physical Books Limiting Reactant Review Sheet Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reactant Review Sheet Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Limiting Reactant Review Sheet Key
 - Setting Reading Goals Limiting Reactant Review Sheet Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Limiting Reactant Review Sheet Key
 - Fact-Checking eBook Content of Limiting Reactant Review Sheet Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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