

Invest 6.1A: Rate Dependant Factors

INTRODUCTION:

Thiosulfate ions react in acid solution to produce a precipitate of sulfur according to the equation:



The rate of the reaction can be followed by measuring the time taken for a fixed amount of sulfur to be produced. This practical first examines the effect on the rate of the reaction when the concentration of thiosulfate ions is altered at a fixed temperature. Then it examines how the rate varies with temperature for a fixed concentration of thiosulfate ions.

ENVIRONMENTAL CARE: Sodium thiosulfate is known as photographers 'hypo' for fixing developed films and prints. To minimize pollution the concentrations of thiosulfate ions have been kept low and only very small amounts of sulfur dioxide are evolved. The residues can be disposed of down the sink.

SAFETY: There are no particular hazards associated with this practical except that you should avoid breathing in any sulfur dioxide that is evolved while you are waiting for the cross to disappear.

PROCEDURE:

1. Concentration.

- Measure 50 cm³ of 0.20 mol dm⁻³ sodium thiosulfate solution into a conical flask.
- Place the flask onto a white piece of paper or note card marked with a black "X".
- Measure out 5 cm³ of 2.0 mol dm⁻³ hydrochloric acid into a measuring cylinder and then add the acid to the thiosulfate starting a stop-clock at the same time.
- Swirl the contents of the flask initially to mix and then allow the flask to remain still on the tile.
- Look down vertically through the solution and time how long it takes for the cross to "just" disappear.
- Rinse the flask out immediately but take care not to get the X on the paper wet.
- Then put 40 cm³ of the 0.20 mol dm⁻³ sodium thiosulfate solution into the flask and add 10 cm³ of water.
- Add 5 cm³ of the hydrochloric acid and time as before.
- Repeat the experiment four more times using 30, 20, 15 and 10 cm³ of the thiosulfate solution making the total volume up to 50 cm³ with water each time.

2. Temperature.

- Measure 10 cm³ of 0.20 mol dm⁻³ thiosulfate solution. Add distilled water to a total volume of 50 cm³ into the conical flask and warm until the temperature is one or two degrees above 30-35 °C on a hot plate.
- Place the flask on the tile marked with a cross and add 5 cm³ of 2.0 mol dm⁻³ hydrochloric acid, timing as before.
- Record the temperature of the mixture after the acid has been added.
- Repeat the experiment using fresh portions of the 0.20 mol dm⁻³ thiosulfate solution each time at temperatures of approximately 40, 50 and 60 °C.

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Table of Contents Lab Investigation 33a

1. Understanding the eBook Lab Investigation 33a
 - The Rise of Digital Reading Lab Investigation 33a
 - Advantages of eBooks Over Traditional Books
2. Identifying Lab Investigation 33a
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lab Investigation 33a
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lab Investigation 33a
 - Personalized Recommendations
 - Lab Investigation 33a User Reviews and Ratings
 - Lab Investigation 33a and Bestseller Lists
5. Accessing Lab Investigation 33a Free and Paid eBooks
 - Lab Investigation 33a Public Domain eBooks
 - Lab Investigation 33a eBook Subscription Services
 - Lab Investigation 33a Budget-Friendly Options

-
6. Navigating Lab Investigation 33a eBook Formats
 - ePub, PDF, MOBI, and More
 - Lab Investigation 33a Compatibility with Devices
 - Lab Investigation 33a Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lab Investigation 33a
 - Highlighting and Note-Taking Lab Investigation 33a
 - Interactive Elements Lab Investigation 33a
 8. Staying Engaged with Lab Investigation 33a
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab Investigation 33a
 9. Balancing eBooks and Physical Books Lab Investigation 33a
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab Investigation 33a
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Lab Investigation 33a
 - Setting Reading Goals Lab Investigation 33a
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Lab Investigation 33a
 - Fact-Checking eBook Content of Lab Investigation 33a
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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