

Graphing Periodic Trends Lab

Name _____ Hr _____

PURPOSE: To determine how certain properties are periodic when the elements are arranged in periods or groups.

DISCUSSION: There are many periodic trends that occur as you cross a period or go down a group/family on the Periodic Table. Graphing various elemental properties will help to illustrate these trends.

REFERENCE:

Angstrom - A unit of distance equal to one ten-billionth of a meter.

Atomic Radius - The distance from the center of the nucleus to the valence electron energy level.

Ionization energy - The energy required to remove an electron from an atom.

DATA

Properties of the First 18 Elements

Element	Atomic Number	Atomic Radius (Angstroms)	Ionization Energy
H	1	0.79	314
He	2	0.49	567
Li	3	2.05	124
Be	4	1.4	215
B	5	1.17	191
C	6	0.91	260
N	7	0.75	335
O	8	0.65	314
F	9	0.57	402
Ne	10	?	?
Na	11	2.23	119
Mg	12	1.72	176
Al	13	1.82	138
Si	14	1.46	168
P	15	1.23	242
S	16	1.09	239
Cl	17	0.97	299
Ar	18	?	?

Using the data above, make a line graph the following data. Draw a dark vertical line to separate the 3 periods. Give each graph a title and label both the x-axis and the y-axis. Connect your points. Predict the plot points for elements with question marks.

1. Atomic Radius (y) vs. Atomic Number (x) for the first 3 periods (18 elements – 3 graphs)
2. Ionization energy (y) vs. Atomic Number (x) for the first 3 periods (18 elements – 3 graphs)

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Table of Contents Lab Graphing Periodic Trends Answers

1. Understanding the eBook Lab Graphing Periodic Trends Answers
 - The Rise of Digital Reading Lab Graphing Periodic Trends Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Lab Graphing Periodic Trends Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lab Graphing Periodic Trends Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lab Graphing Periodic Trends Answers
 - Personalized Recommendations
 - Lab Graphing Periodic Trends Answers User Reviews and Ratings
 - Lab Graphing Periodic Trends Answers and Bestseller Lists

5. Accessing Lab Graphing Periodic Trends Answers Free and Paid eBooks
 - Lab Graphing Periodic Trends Answers Public Domain eBooks
 - Lab Graphing Periodic Trends Answers eBook Subscription Services
 - Lab Graphing Periodic Trends Answers Budget-Friendly Options
6. Navigating Lab Graphing Periodic Trends Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Lab Graphing Periodic Trends Answers Compatibility with Devices
 - Lab Graphing Periodic Trends Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lab Graphing Periodic Trends Answers
 - Highlighting and Note-Taking Lab Graphing Periodic Trends Answers
 - Interactive Elements Lab Graphing Periodic Trends Answers
8. Staying Engaged with Lab Graphing Periodic Trends Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab Graphing Periodic Trends Answers
9. Balancing eBooks and Physical Books Lab Graphing Periodic Trends Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab Graphing Periodic Trends Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Lab Graphing Periodic Trends Answers
 - Setting Reading Goals Lab Graphing Periodic Trends Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Lab Graphing Periodic Trends Answers
 - Fact-Checking eBook Content of Lab Graphing Periodic Trends Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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