

Lab 10 – pH and Buffers

A. Objectives:

1. Constructive proper titration and titration.
2. Construct and explain how the pH scale works.
3. Construct and explain the concept of buffer capacity.
4. Construct acids and bases with strong acids.
5. Calculate the pH of various buffer solutions.
6. Calculate equilibrium pH, buffer capacity and buffer capacity of a buffer.

B. Background: Titration, Buffer Capacity and Buffer

1. Titration

Titration is a chemical reaction between two substances. The reaction involves the transfer of electrons from one substance to another. Titration is a process of determining the concentration of a solution by reacting it with a known volume of a standard solution.

2. Buffering and titration. A buffer is a solution that resists changes in pH. It is a solution that contains a weak acid and its conjugate base. The buffering capacity of a buffer is the amount of acid or base that can be added to the solution without causing a significant change in pH. The buffering capacity of a buffer is determined by the concentration of the weak acid and its conjugate base.

3. Buffering and titration. A buffer is a solution that resists changes in pH. It is a solution that contains a weak acid and its conjugate base. The buffering capacity of a buffer is the amount of acid or base that can be added to the solution without causing a significant change in pH. The buffering capacity of a buffer is determined by the concentration of the weak acid and its conjugate base.

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Concentration (molarity)	pH	Concentration	pH	Concentration	pH
0.1 M acetic acid		0.1 M acetic acid		0.1 M acetic acid	
0.1 M acetic acid		0.1 M acetic acid		0.1 M acetic acid	
0.1 M acetic acid		0.1 M acetic acid		0.1 M acetic acid	

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