

Experiment 14

Heat Effects and Calorimetry

Heat is a form of energy, sometimes called thermal energy, that can pass spontaneously from an object at a high temperature to an object at a lower temperature. If the two objects are in contact, they will, given sufficient time, both reach the same temperature.

Heat flow is ordinarily measured in a device called a calorimeter. A calorimeter is simply a container with insulating walls, made so that essentially no heat is exchanged between the contents of the calorimeter and the surroundings. Within the calorimeter chemical reactions may occur or heat may pass from one part of the contents to another, but no heat flows into or out of the calorimeter from or to the surroundings.

A. Specific Heat

When heat flows into a substance, the temperature of that substance will increase. The quantity of heat q required to cause a temperature change Δt of any substance is proportional to the mass m of the substance and the temperature change, as shown in Equation 1. The proportionality constant is called the specific heat, $S.H.$, of that substance.

$$q = (\text{specific heat}) \times m \times \Delta t = S.H. \times m \times \Delta t \quad (1)$$

The specific heat can be considered to be the amount of heat required to raise the temperature of one gram of the substance by 1°C (if you make m and Δt in Equation 1 both equal to 1, then q will equal $S.H.$). Amounts of heat are measured in either joules or calories. To raise the temperature of 1 g of water by 1°C , 4.18 joules of heat must be added to the water. The specific heat of water is therefore 4.18 joules/g $^\circ\text{C}$. Since 4.18 joules equals 1 calorie, we can also say that the specific heat of water is 1 calorie/g $^\circ\text{C}$. Ordinarily heat flow into or out of a substance is determined by the effect that that flow has on a known amount of water. Because water plays such an important role in these measurements, the calorie, which was the unit of heat most commonly used until recently, was actually defined to be equal to the specific heat of water.

The specific heat of a metal can readily be measured in a calorimeter. A weighed amount of metal is heated to some known temperature and is then quickly poured into a calorimeter that contains a measured amount of water at a known temperature. Heat flows from the metal to the water, and the two equilibrate at some temperature between the initial temperatures of the metal and the water.

Assuming that no heat is lost from the calorimeter to the surroundings, and that a negligible amount of heat is absorbed by the calorimeter walls, the amount of heat that flows from the metal as it cools is equal to the amount of heat absorbed by the water.

In thermodynamic terms, the heat flow for the metal is equal in magnitude but opposite in direction, and hence in sign, to that for the water. For the heat flow q ,

$$q_{\text{H}_2\text{O}} = -q_{\text{metal}} \quad (2)$$

If we now express heat flow in terms of Equation 1 for both the water and the metal M , we get

$$q_{\text{H}_2\text{O}} = S.H._{\text{H}_2\text{O}} m_{\text{H}_2\text{O}} \Delta t_{\text{H}_2\text{O}} = -S.H._M m_M \Delta t_M \quad (3)$$

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Kathleen Armour



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Laboratory Manual for World of Chemistry Melvin D. Joesten, David O. Johnston, John T. Netterville, James L. Wood, 1987 Thermochemistry and Its Applications to Chemical and Biochemical Systems M.A.V. Ribeiro Da Silva, 2012-12-06 Proceedings of the NATO Advanced Study Institute on Thermochemistry Today and Its Role in the Immediate Future Viano do Castelo Portugal July 5 15 1982 Laboratory Manual for Chemistry, Man, and Society Mark Martin Jones, John W. Dawson, 1972 A Laboratory Manual of Analytical Methods of Protein Chemistry, Including Polypeptides Peter Alexander, Richard Joseph Block, Harold Palmer Lundgren, 1960 In the last fifteen years there has been a revolution in the techniques available for the analysis and isolation of proteins Every time a new technique has been introduced numerous papers have appeared describing modifications to it and the research worker who wishes to employ these methods is faced with a very serious problem in deciding which particular variant to use These volumes are intended to provide the fullest practical detail so that any scientist can follow the procedure by using this book alone and without having recourse to the original literature The techniques which are described in full are ones in which all the authors have had first hand experience and the descriptions contain those small but important points which save so much time In the first volume separation and isolation procedures are discussed the second concerns its analysis and reactivity and the third volume with the measurement of the macromolecular properties of proteins **Laboratory Manual for Chemistry** Karen Timberlake, 1979 **Physical Experiments** John Francis Woodhull, May Belle Van Arsdale, 1900 Free Night School of Engineering Cooper Union for the Advancement of Science and Art. School of Engineering, 1928 Electrochemical and Metallurgical Industry Eugene Franz Roeber, Howard Coon Parmelee, 1909 A Laboratory Manual of Inorganic Chemistry, with an Introduction to Qualitative Analysis John Bernard Ekeley, 1934 **Host Bibliographic Record for Boundwith Item Barcode 30112114006288 and Others**, 1901 **Publications of the National Institute of Standards and Technology ... Catalog** National Institute of Standards and Technology (U.S.), 1983 *NBS Special Publication*, 1968 **Publications of the National Bureau of Standards** United States. National Bureau of Standards, 1982 *Joint Volumes of Papers Presented to the Legislative Council and Legislative Assembly* New South Wales. Parliament, 1906 Includes various departmental reports and reports of commissions Cf Gregory Serial publications of foreign governments 1815 1931 Calendar - McGill University McGill University, 1909 Chemical Engineering Catalog, 1919 Guide for Safety in the Chemical Laboratory Manufacturing Chemists' Association (U.S.), 1972 Annual Calendar of McGill College and University, Montreal McGill University, 1909 **Report** Cooper Union for the Advancement of Science and Art, 1904 Chemical & Metallurgical Engineering Eugene Franz Roeber, Howard Coon Parmelee, 1928

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