

# Sulfur clock lab

## Experimental procedure

1. Measurement of the formation of visible product
2. Measurement of the initial concentration of reactants
3. Prepare solutions A, B, C, D and E as follows:

see Table 10.10.10.10

	Reactants	Reagents	Apparatus
A	1.0 M	0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{K}_2\text{S}_2\text{O}_8$	Reaction setup
B	0.5 M	0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{K}_2\text{S}_2\text{O}_8$	0.1 M, 0.01 M, 0.01 M
C	0.25 M	0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{K}_2\text{S}_2\text{O}_8$	0.1 M, 0.01 M, 0.01 M
D	0.125 M	0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{K}_2\text{S}_2\text{O}_8$	0.1 M, 0.01 M, 0.01 M

4. Add the solution using a graduated cylinder with completely standard chemical formulae for the reaction.
5. Transfer exactly the calculated amount of  $\text{H}_2\text{O}_2$  solution to the reaction mixture using a graduated cylinder.
6. Add exactly 1.00 mL of 0.1 M  $\text{K}_2\text{S}_2\text{O}_8$  solution to the reaction mixture.
7. Prepare and use the standard solutions.
8. Prepare solutions A, B, C, D and E as follows:
  - a. Transfer 10.0 mL of 0.1 M  $\text{H}_2\text{O}_2$  solution to a beaker.
  - b. Add 1.0 mL of 0.1 M  $\text{K}_2\text{S}_2\text{O}_8$  solution using a graduated cylinder.
  - c. Dilute to the 100 mL mark using distilled water.
  - d. Use the same procedure for all other solutions.
9. Transfer solution A into the graduated cylinder as 100 mL sample. Use the same method for all other solutions.

# Lab Sulfur Clock Answers

**C Cleary**



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**The Chemical News and Journal of Physical Science** ,1903      **Chemical News and Journal of Physical Science** ,1902      **Interior Department Appropriation Bill for 1945** United States. Congress. House. Committee on Appropriations. Subcommittee on Interior Department,1944      **South Dakota Farm & Home Research** ,1994      *The Journal of Industrial and Engineering Chemistry* ,1917      **Chemical News and Journal of Industrial Science** ,1902

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