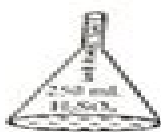


অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার

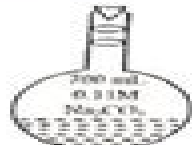


পটাসিয়াম কপ্লো, এখানে মোল পটাসিয়াম, আরও কপ্লোয়ন শীর্ষস্থানীয় কপ্লোয়নমুকের পটাসিয়াম এবং ময়েল টেম্পের পটাসিয়াম ল্যাবরেটরির অধ্যয়নিক সেএক্স হলেখে। এপুলের অনুবীলন করলে তুমি এ অধ্যয়নিক থেকে সেএক্সের ল্যাবরেটরির অধ্যয়নিক সেএক্স হলেখে।

প্রশ্ন ১



পাত্র-A



পাত্র-B

সি. কে. ১০। এফ. কে. ১০।

- মোলের প্রকণ কী?
- "কেনাস পানিযুক্ত লবণের প্রাকৃত্য, আপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়"— ব্যাখ্যা করো।
- A-পাত্রের বিন্যমান এসিডের ভর কত নির্ণয় করো।
- প্রকণ দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অক্সিজেন ফাঙ্কীয় হবে— বিয়োজন করো।

১ নং প্রশ্নের উত্তর

1000 mL প্রকণে কোন প্রকণের 1 mol প্রবীভূত থাকলে উক্ত প্রকণকে মোলার প্রকণ বলে।

কেনাস পানিযুক্ত লবণের প্রাকৃত্য আপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়। এর কারণ, আপমাত্রা বৃদ্ধির সাথে সাথে প্রাকৃত্য বৃদ্ধি পেতে থাকলেও একটি নির্দিষ্ট আপমাত্রার কেনাস পানি দুই লবণ নিষ্কৃতিত হয়। তখন প্রাকৃত্য হ্রাস পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কথা বলা যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর ক্ষেত্রে পানিতে প্রবীভূত করার সময় বিয়োজন প্রক্রিয়াটি আপমাত্রা বৃদ্ধি 32.4°C আপমাত্রার উপরে $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ যখন Na_2SO_4 এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি আপোৎপাদী হয়। এ কারণে 32.4°C আপমাত্রার উপরে আপমাত্রা বৃদ্ধিজনিত কারণে প্রাকৃত্য হ্রাস ঘটে।

২

আমরা জানি,
 $w = SMV$
 $= (0.5 \times 98 \times 0.25) \text{ g}$
 $= 12.25 \text{ g}$

অতএব, প্রবীভূত H_2SO_4 এর ভর
 $= 12.25 \text{ g}$

উদীপকের A-পাত্রের অবস্থিতি,
 H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$

প্রকণের ঘনমাত্রা,
 $S = 0.5 \text{ M}$
 প্রকণের আয়তন,
 $V = 250 \text{ mL}$
 $= \frac{250}{1000} = 0.25 \text{ L}$

প্রবীভূত H_2SO_4 এর ভর, $w = ?$

প্রশ্ন-A: H_2SO_4 এর মোল সংখ্যা = $\left(0.5 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.125 \text{ mol}$

প্রশ্ন-B: Na_2CO_3 এর মোল সংখ্যা = $\left(0.11 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.077 \text{ mol}$

সম্বন্ধিত বিক্রিয়া: $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$

সমীকরণ হতে পাই, 1 mol $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$

1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1 mol এর সাথে

অবশিষ্ট H_2SO_4 এর পরিমাণ = $(0.125 - 0.077) \text{ mol}$
 $= 0.048 \text{ mol}$

পাত্র-A এর প্রকণের সাথে পাত্র-B এর প্রকণের বিক্রিয়ার ফলে 0.048 mol H_2SO_4 অবশিষ্ট থাকায় মিশ্রিত প্রকণের প্রকৃতি হবে অম্লীয়।

প্রশ্ন ২ এককাল ছাত্র ল্যাবরেটরিতে X-পদ্ধতিতে 10 mg এর প্রকণ পরিমাণ এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2% $\frac{W}{V}$ HCl এর দ্রব প্রকণ করে।

সি. কে. ১০। এফ. কে. ১০।

- নোড কাকে বলে?
- Cl_2 অপোলার— ব্যাখ্যা করো।
- উদীপকের এসিডটির মোলার ঘনমাত্রা নির্ণয় করো।
- উদীপকের কোন পদ্ধতি পরিবেশবান্ধব? বিয়োজন করো।

২ নং প্রশ্নের উত্তর

নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেক্ট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা বলা হয় অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন গঠিত হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোলার হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও H এর তড়িৎ ঋণাত্মকতা 2.1, তাই Cl_2 বন্ধন সমযোজী ও অপোলার যৌগ। উল্লেখ্য যৌলিক অণুর ক্ষেত্রে এহুণ ঘাটতে পারে।

উদীপকে বর্ণিত এসিডটি হলো 2% $\left(\frac{W}{V}\right) \text{ HCl}$

2% $\left(\frac{W}{V}\right) \text{ HCl}$ এর অর্থ হলো—

100 mL প্রকণে HCl এর ভর 2g

∴ 1000 mL " " " " $\left(\frac{2 \times 1000}{100}\right) \text{ g}$
 $= 20 \text{ g}$

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1. Understanding the eBook Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper
 - The Rise of Digital Reading Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper
 - Personalized Recommendations

- Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper User Reviews and Ratings
- Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper and Bestseller Lists
- 5. Accessing Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper Free and Paid eBooks
 - Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper Public Domain eBooks
 - Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper eBook Subscription Services
 - Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper Budget-Friendly Options
- 6. Navigating Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper Compatibility with Devices
 - Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper
 - Highlighting and Note-Taking Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper
 - Interactive Elements Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper
- 8. Staying Engaged with Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper
- 9. Balancing eBooks and Physical Books Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper
 - Setting Reading Goals Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper
 - Fact-Checking eBook Content of Mcq Answer Of Hsc Maharashtra 2015 Chemistry Paper

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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