

3.1**MEASUREMENTS AND THEIR UNCERTAINTY****Section Review****Objectives**

- Convert measurements to scientific notation
- Distinguish among the accuracy, precision, and error of a measurement
- Identify the number of significant figures in a measurement and in the result of a calculation

Vocabulary

- | | | |
|-----------------------|----------------------|-----------------------|
| • measurement | • precision | • error |
| • scientific notation | • accepted value | • percent error |
| • accuracy | • experimental value | • significant figures |

Key Equations

- Error = experimental value - accepted value
- Percent error = $\frac{|\text{error}|}{\text{accepted value}} \times 100\%$

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

- The 1 of a measurement describes how close the measurement comes to the true value. The 2 of a measurement depends on its reproducibility. An 3 is a value measured in the lab. 4 is calculated by subtracting the 5 from an experimental value. Percent error is calculated by dividing the 6 of the error by the accepted value and then multiplying by 7.
- Large and small numbers are more easily handled when expressed in 8. Significant figures in a measurement include all of the digits that are 9 plus a last digit that is 10.
- | |
|-----------|
| 1. _____ |
| 2. _____ |
| 3. _____ |
| 4. _____ |
| 5. _____ |
| 6. _____ |
| 7. _____ |
| 8. _____ |
| 9. _____ |
| 10. _____ |

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web k sp of calcium hydroxide 1 author j m mccormick last update ocober 13 2013 introduction an equilibrium constant K is related to ΔG for a process through eqn 1 and as long as ΔH and ΔS are independent of temperature eqn 1 can be combined with eqn 2 to determine ΔH and ΔS from the temperature dependence of K

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web your primary objective in this experiment is to test a saturated solution of calcium hydroxide and use your observations and measurements to calculate the ksp of the compound you will do this by titrating the prepared Ca(OH)_2 solution with a standard hydrochloric acid solution

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web in today s experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the

concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2(\text{s}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq})$ the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is used as a pH regulating agent notice that calcium hydroxide is divalent and thus has twice

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution $\text{Ca(OH)}_2 \rightarrow \text{Ca}^{2+} + 2\text{OH}^{-}$ $5 \times 10^{-2} \text{ M}$ calcium ions $5 \times 10^{-2} \text{ M}$ 0 M hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the ksp of a slightly soluble basic salt and how solubility changes with temperature ksp is another form of equilibrium expression that deals with solubility the reaction that will take place is $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^{-}$

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web the chemical equation used was $\text{Ca(OH)}_2(\text{s}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq})$ the solubility product was determined to be ksp $\text{Ca}^{2+}\text{OH}^{-}$ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

18.1 solubility product constant ksp chemistry libretexts - Feb 01 2023

web jul 12 2023 because the concentration of a pure solid such as Ca_3PO_4 is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $\text{Ca}_3\text{PO}_4(\text{s}) \rightleftharpoons 3\text{Ca}^{2+}(\text{aq}) + 4\text{PO}_4^{3-}(\text{aq})$ Ksp Ca_3PO_4

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web lab report 8 determinin of calcium hydroxide the purpose of this was to titrat e a saturated Ca(OH)_2 soultion with a st andard hcl solution and find the ksp or the oh concentration was f ound and then the concentration was found and u for this experiment a titration of calcium hydroxide with hydrochlo

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web question determining the ksp of calciumhydroxide Ca(OH)_2 s Ca^{2+} aq 2OH^- aq for this experiment the titration of calciumhydroxide with hydrochloric acid wasinvestigated determining the ksp of calcium hydroxide Ca(OH)_2 s Ca^{2+} aq 2OH^- aq for this experiment the titration of calcium

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for Ca was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

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web we will filter the solid calcium hydroxide from the solution and then titrate the dissolved calcium hydroxide in solution with a known concentration of hydrochloric acid using titrimetric calculations we will determine the Ca^{2+} and OH^- concentrations in solution

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web calculate the ksp for calcium hydroxide $\text{ksp Ca}^{2+} \text{OH}^-$ 2 0 0 2 $\text{ksp} = 5 \times 10^{-4}$ find the accepted value of the ksp for calcium

hydroxide and compare it with your value discuss the discrepancy and suggest possible sources of experimental error
accepted value for K_{sp} for sodium hydroxide is 5×10^{-6}

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