

BUILD Vocabulary

ionic bond Chemical bond formed when one or more electrons are transferred from one atom to another.

ion atom that has a positive or negative charge.

covalent bond type of bond between atoms in which the electrons are shared.

molecule smallest unit of most compounds that displays all the properties of that compound.

van der Waals forces weak attraction that develops between oppositely charged regions of nearby molecules.

Preface The prefix *co-* and *co-* mean "together." A compound is made up of elements joined together, and a covalent bond connects atoms that share electrons together. **How are covalent bonds different from ionic bonds?**

Chemical Bonds

KEY QUESTION What are the main types of chemical bonds?

The atoms in compounds are held together by chemical bonds. Bonds are formed by the electrons in the outer shell, which are called valence electrons. The main types of chemical bonds are ionic bonds and covalent bonds.

Ionic Bonds **Ionic bonds** are formed when one or more electrons are transferred from one atom to another. This transfer causes the atom losing the electrons to become positively charged. The atom gaining the electrons becomes negatively charged. The positively and negatively charged atoms become **ions**. The oppositely charged ions are attracted to each other and form an ionic bond.

Covalent Bonds Sometimes atoms share electrons rather than transferring them. These shared electrons move around the nuclei of both atoms. The bond formed by sharing electrons is called a **covalent bond**. Atoms can share between two and six electrons. The structure that results is known as a **molecule**. A molecule is the smallest unit of most compounds.

Weak Interactions The strongest chemical bonds are ionic or covalent. Weak interactions between atoms are also important. Within living organisms, many molecules interact only briefly. **Van der Waals forces** are one type of weak interaction between molecules. They occur when molecules are very close together.

Hydrogen bonds are another type of weak interaction. They occur between a hydrogen atom of one molecule and an oxygen or nitrogen atom of a neighboring molecule.

Visual Reading Skill Covalent Bonds

1. What kind of chemical bond holds the water molecule together? _____
2. What particles are shared between the atoms in the molecule? _____
3. What is the other main type of chemical bond? _____
4. What is the difference between covalent and ionic bonds? _____



Water molecule (H_2O)

Properties of Water

READING TOOL Cause and Effect As you read the part of the lesson under "The Water Molecule," use the table below to list the causes and effects of the properties of water. Many of the causes may be the same, but each characteristic will have a different effect. Part of the table is completed for you.

Characteristic of Water	Cause	Effect
Hydrogen Bonding		The slightly negative end of a water molecule is attracted to the slightly positive end of another water molecule.
Expands Upon Freezing		
Dissolves Many Substances	Water forms hydrogen bonds.	
Cohesion		
Adhesion		
Heat Capacity		

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healthcare and infection control professionals as well as students for preventing and controlling infection Biofilms are the most common mode of bacterial growth in nature Highly resistant to antibiotics and antimicrobials biofilms are the source of more than 65 percent of health care associated infections HCAI which according to the WHO affect 1 4 million people annually Biofilms are involved in 80 percent of all microbial infections in the body including those associated with medical devices such as catheters endotracheal tubes joint prostheses and heart valves Biofilms are also the principle causes of infections of the middle ear dental caries gingivitis prostatitis and cystic fibrosis Importantly biofilms also significantly delay wound healing and reduce antimicrobial efficiency in at risk or infected skin wounds Provides specific procedures for controlling and preventing infection Includes case studies of HCAI and identifies appropriate treatments Presents national government standards for infection prevention and control Includes extensive references and links to websites for further information

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Fundamentals of Space Life Sciences Susanne Elaine Churchill,1997 *Subject Guide to Books in Print* ,1997 *The Role of the Study Director in Nonclinical Studies* William J. Brock,Barbara Mounho,Lijie Fu,2014-06-03 A single source reference with a broad and holistic overview of nonclinical studies this book offers critical training material and describes regulations of nonclinical testing through guidelines models case studies practical examples and worldwide perspectives The book Provides a complete overview of nonclinical study organization conduct and reporting and describes the roles and responsibilities of a Study Director to manage an effective study Covers regulatory and scientific concepts including international testing and Good Laboratory Practice GLP compliance with guidelines and animal models Features a concluding chapter that compiles case studies lessons learned from those that have served as a Study Director for many years Addresses the entire spectrum of nonclinical testing making it applicable to those in the government laboratories and those actively involved in in all sectors of industry

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