

$$\text{percent yield} = \frac{\text{actual yield (g)}}{\text{theoretical yield (g)}} \times 100\%$$

↳ from B-E-A chart

Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation.
- Identify the given (with units) and what you want to find (with units).
- Show set up with units. Give final answer with units and label.

1. Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



1 mol H_2O : 18.02 g/mol

36 g $\frac{1\text{ mol H}_2\text{O}}{18.02\text{ g}} = 1.99\text{ mol H}_2\text{O}$

1 mol H_2 : 2.02 g

2 mol $\text{H}_2 = \boxed{4.04\text{ g H}_2}$

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



1 mol $\text{NaCl} = 58.44\text{ g}$
 $581\text{ g NaCl} \times \frac{1\text{ mol NaCl}}{58.44\text{ g}} = 9.94\text{ mol NaCl}$

1 mol $\text{Na} = 22.99\text{ g}$

$9.94\text{ mol Na} \times 22.99\text{ g} = \boxed{228.52\text{ g Na}}$

Modeling Chemistry U8 Empfrm Practice V2

G Orfield



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Table of Contents Modeling Chemistry U8 Empfrm Practice V2

1. Understanding the eBook Modeling Chemistry U8 Empfrm Practice V2
 - The Rise of Digital Reading Modeling Chemistry U8 Empfrm Practice V2
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry U8 Empfrm Practice V2
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry U8 Empfrm Practice V2
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry U8 Empfrm Practice V2
 - Personalized Recommendations

- Modeling Chemistry U8 Empfrm Practice V2 User Reviews and Ratings
- Modeling Chemistry U8 Empfrm Practice V2 and Bestseller Lists
- 5. Accessing Modeling Chemistry U8 Empfrm Practice V2 Free and Paid eBooks
 - Modeling Chemistry U8 Empfrm Practice V2 Public Domain eBooks
 - Modeling Chemistry U8 Empfrm Practice V2 eBook Subscription Services
 - Modeling Chemistry U8 Empfrm Practice V2 Budget-Friendly Options
- 6. Navigating Modeling Chemistry U8 Empfrm Practice V2 eBook Formats
 - ePub, PDF, MOBI, and More
 - Modeling Chemistry U8 Empfrm Practice V2 Compatibility with Devices
 - Modeling Chemistry U8 Empfrm Practice V2 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry U8 Empfrm Practice V2
 - Highlighting and Note-Taking Modeling Chemistry U8 Empfrm Practice V2
 - Interactive Elements Modeling Chemistry U8 Empfrm Practice V2
- 8. Staying Engaged with Modeling Chemistry U8 Empfrm Practice V2
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry U8 Empfrm Practice V2
- 9. Balancing eBooks and Physical Books Modeling Chemistry U8 Empfrm Practice V2
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry U8 Empfrm Practice V2
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry U8 Empfrm Practice V2
 - Setting Reading Goals Modeling Chemistry U8 Empfrm Practice V2
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry U8 Empfrm Practice V2
 - Fact-Checking eBook Content of Modeling Chemistry U8 Empfrm Practice V2

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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