

Unit 7 Chemical Reactions
PRACTICE TEST/ REVIEW

- _____ 1. A chemical reaction is a process in which
- a. products change into reactants
 - b. the law of conservation of mass applies
 - c. substances can change state
 - d. all of these
- _____ 2. During a chemical reaction,
- a. new elements are produced
 - b. atoms are rearranged
 - c. atoms are destroyed
 - d. elements are destroyed
- _____ 3. An equation is balanced by
- a. changing subscripts
 - b. adding coefficients
 - c. erasing elements as necessary
 - d. adding elements as necessary
- _____ 4. What are the reactants in the following chemical equation:
- $$\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$$
- a. zinc and copper
 - b. zinc sulfate and copper
 - c. zinc and copper (II) sulfate
 - d. only zinc
- _____ 5. What are the products in the above equation?
- a. zinc and copper
 - b. zinc sulfate and copper
 - c. zinc and copper (II) sulfate
 - d. only zinc
- _____ 6. The breaking down of a substance into two or simpler substances is _____
- a. Displacement
 - b. Synthesis
 - c. Decomposition
 - d. Double displacement
- _____ 7. Each substance to the right of the arrow in a chemical equation is a(n)
- a. Reactant
 - b. Catalyst
 - c. Product
 - d. Inhibitor
- _____ 8. According to the law of conservation of mass, if two atoms of hydrogen are used as a reactant, how many atoms of hydrogen must be part of the product?
- a. 4
 - b. 2
 - c. 1
 - d. 0
- _____ 9. What type of reaction is shown in the following chemical equation, $\text{NH}_4\text{Cl} \rightarrow \text{NH}_3 + \text{HCl}$?
- a. Displacement
 - b. Synthesis
 - c. Decomposition
 - d. Double displacement

Modeling Chemistry Unit 7 Chemical Reactions Answer Key

Michal Rosen-Zvi



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Table of Contents Modeling Chemistry Unit 7 Chemical Reactions Answer Key

1. Understanding the eBook Modeling Chemistry Unit 7 Chemical Reactions Answer Key
 - The Rise of Digital Reading Modeling Chemistry Unit 7 Chemical Reactions Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit 7 Chemical Reactions Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry Unit 7 Chemical Reactions Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit 7 Chemical Reactions Answer Key
 - Personalized Recommendations
 - Modeling Chemistry Unit 7 Chemical Reactions Answer Key User Reviews and Ratings
 - Modeling Chemistry Unit 7 Chemical Reactions Answer Key and Bestseller Lists
5. Accessing Modeling Chemistry Unit 7 Chemical Reactions Answer Key Free and Paid eBooks

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 - ePub, PDF, MOBI, and More
 - Modeling Chemistry Unit 7 Chemical Reactions Answer Key Compatibility with Devices
 - Modeling Chemistry Unit 7 Chemical Reactions Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 7 Chemical Reactions Answer Key
 - Highlighting and Note-Taking Modeling Chemistry Unit 7 Chemical Reactions Answer Key
 - Interactive Elements Modeling Chemistry Unit 7 Chemical Reactions Answer Key
- 8. Staying Engaged with Modeling Chemistry Unit 7 Chemical Reactions Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry Unit 7 Chemical Reactions Answer Key
- 9. Balancing eBooks and Physical Books Modeling Chemistry Unit 7 Chemical Reactions Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry Unit 7 Chemical Reactions Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry Unit 7 Chemical Reactions Answer Key
 - Setting Reading Goals Modeling Chemistry Unit 7 Chemical Reactions Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry Unit 7 Chemical Reactions Answer Key
 - Fact-Checking eBook Content of Modeling Chemistry Unit 7 Chemical Reactions Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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