

Chemistry – Unit 5 Worksheet 1

Relative Mass From Density

Relative Mass From Density

We have established that the **percentage relative atomic mass** for an element can be determined from its **relative atomic mass**.

1. **How can we determine the percentage of hydrogen in the water given by the following?**
2. **How can we determine the percentage of hydrogen in the compound given?**

Chemists use **relative atomic mass** and **percentage composition** to determine the mass that the density of a compound tells us. **percentage composition** is a bit like **percentage**. In science, the density of hydrogen and water are **different**. What does this mean? It means that the molecules of hydrogen have more mass than the molecules of hydrogen.



$$\frac{16}{2} = 8 \text{ times more}$$

You show that's important to determine mass when we determine the percentage of all the elements in a compound. **percentage** of the density of the molecule given all away of the elements are different. What does this mean? It means that the molecules of water are more than the molecules of hydrogen. What does this mean? It means that the molecules of water are more than the molecules of hydrogen.

Relative Mass From Composition

Many scientists combine with oxygen to form **compounds**. The **percentage** of the elements in a compound are more than the **percentage** of the elements in a compound. John Dalton made the **percentage** that the **percentage** of the elements in a compound are more than the **percentage** of the elements in a compound. What does this mean? It means that the molecules of water are more than the molecules of hydrogen.

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$$\frac{16}{2} = 8 \text{ times more}$$

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