

CHE1APL – Environmental Chemistry Workshop 1

This workshop sheet consists of two sections. Section 1 includes questions for which your tutor will be available to check your workings and help you during the workshop session. To aid your learning, make sure that you attempt the workshop questions before coming to class. Section 2 includes short-answer or multiple-choice questions that you can work through in your own time; answers to which will be provided during the workshop.

Section 1.

In this section, you will use some basic mathematics skills that are essential in science to manipulate logarithms and exponential functions, as well as solve simple equations using basic calculus. You can refresh your maths skills by examining the content of Appendix I on "Mathematics for Chemistry: Exponential Notation, Logarithms, Graphing, and Quadratic Equations" in the prescribed textbook. Additional on-line maths help is available from OWL exercises via the LMS for CHE1APL.

Factors Affecting Solubility: Pressure and Temperature (Section 12.4)

1. The Henry's law constant for N_2 in water at 37°C is $6.2 \times 10^{-6} \text{ mol L}^{-1} \text{ kPa}^{-1}$. Calculate the solubility (in mol L^{-1}) of N_2 in water, when the partial pressure of N_2 is 80.7 kPa?
A. $1.4 \times 10^{-9} \text{ mol L}^{-1}$ B. $6.5 \times 10^{-7} \text{ mol L}^{-1}$ C. $5.0 \times 10^{-4} \text{ mol L}^{-1}$ D. $2.0 \times 10^{-3} \text{ mol L}^{-1}$

Henry's Law: $s = k_H \times p$

solubility (N_2) = Henry's constant $\times$ partial pressure (N_2)

$$s = 6.2 \times 10^{-6} \text{ mol L}^{-1} \text{ kPa}^{-1} \times 80.7 \text{ kPa}$$

$$s = 5 \times 10^{-4} \text{ mol L}^{-1}$$

Colligative Properties: Osmosis (Section 12.6)

2. A solution is prepared by dissolving 4.78 g of an unknown non-electrolyte in enough water to make 375 mL of solution. The osmotic pressure of the solution is 135 kPa at 27°C . Determine the molar mass of the solute? ($R = 8.314 \text{ L kPa K}^{-1} \text{ mol}^{-1}$)
A. 21.2 g mol^{-1} B. 49.4 g mol^{-1} C. 96.8 g mol^{-1} D. 236 g mol^{-1}

Use osmotic pressure equation.

$$\Pi = icRT$$

Where for a non-electrolyte (such as sucrose) $i=1$

$$135 \text{ kPa} = 1.00 \times c \text{ mol L}^{-1} \times 8.314 \text{ L kPa K}^{-1} \text{ mol}^{-1} \times 300 \text{ K}$$

$$c = (135 \text{ kPa}) / (8.314 \text{ L kPa K}^{-1} \text{ mol}^{-1} \times 300 \text{ K}) = 0.05413 \text{ mol L}^{-1}$$

$$\text{now, } c = (\text{mass/MM}) / V,$$

$$\text{therefore, MM} = \text{mass} / (c \cdot V) = 4.78 \text{ g} / (0.05413 \times 0.375) = 235.5 \text{ g mol}^{-1} = 236 \text{ g mol}^{-1}$$

3. What is the osmotic pressure of a 0.126 mol L^{-1} solution of sodium sulphate (Na_2SO_4) at 25°C ? Assume the van't Hoff factor is 2.51.

A. 783 kPa B. 65.7 kPa C. 312 kPa D. 936 kPa

Use osmotic pressure equation for an electrolyte

$$\Pi = icRT$$

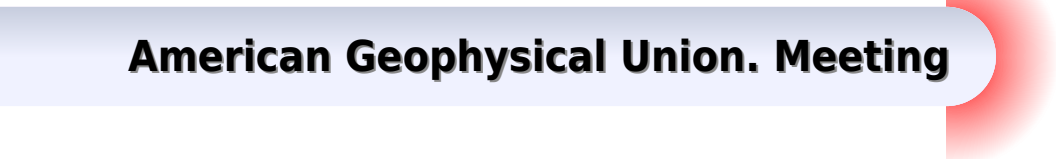
$$\Pi = 2.51 \times 0.126 \text{ mol L}^{-1} \times 8.314 \text{ L kPa K}^{-1} \text{ mol}^{-1} \times 298 \text{ K}$$

$$\Pi = 784 \text{ kPa}$$

Concentration-Time Relationships: Integrated Rate Equations (Section 18.5)

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Table of Contents Modeling Chemistry U7 Ws1 V2 0 Answers

1. Understanding the eBook Modeling Chemistry U7 Ws1 V2 0 Answers
 - The Rise of Digital Reading Modeling Chemistry U7 Ws1 V2 0 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry U7 Ws1 V2 0 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry U7 Ws1 V2 0 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry U7 Ws1 V2 0 Answers
 - Personalized Recommendations
 - Modeling Chemistry U7 Ws1 V2 0 Answers User Reviews and Ratings
 - Modeling Chemistry U7 Ws1 V2 0 Answers and Bestseller Lists

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 - Modeling Chemistry U7 Ws1 V2 0 Answers eBook Subscription Services
 - Modeling Chemistry U7 Ws1 V2 0 Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Modeling Chemistry U7 Ws1 V2 0 Answers Compatibility with Devices
 - Modeling Chemistry U7 Ws1 V2 0 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry U7 Ws1 V2 0 Answers
 - Highlighting and Note-Taking Modeling Chemistry U7 Ws1 V2 0 Answers
 - Interactive Elements Modeling Chemistry U7 Ws1 V2 0 Answers
8. Staying Engaged with Modeling Chemistry U7 Ws1 V2 0 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry U7 Ws1 V2 0 Answers
9. Balancing eBooks and Physical Books Modeling Chemistry U7 Ws1 V2 0 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry U7 Ws1 V2 0 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modeling Chemistry U7 Ws1 V2 0 Answers
 - Setting Reading Goals Modeling Chemistry U7 Ws1 V2 0 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modeling Chemistry U7 Ws1 V2 0 Answers
 - Fact-Checking eBook Content of Modeling Chemistry U7 Ws1 V2 0 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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