

## Section 4.5

# Types of Aqueous Solutions and Solubility

## Solubility Rules for Ionic Compounds in Water

TABLE 4.1 Solubility Rules for Ionic Compounds in Water

Compounds Containing the Following Ions Are Generally Soluble	Exceptions
$\text{Li}^+$ , $\text{Na}^+$ , $\text{K}^+$ , and $\text{NH}_4^+$	None
$\text{NO}_3^-$ and $\text{C}_2\text{H}_3\text{O}_2^-$	None
$\text{Cl}^-$ , $\text{Br}^-$ , and $\text{I}^-$	When these ions pair with $\text{Ag}^+$ , $\text{Hg}_2^{2+}$ , or $\text{Pb}^{2+}$ , the resulting compounds are insoluble.
$\text{SO}_4^{2-}$	When $\text{SO}_4^{2-}$ pairs with $\text{Sr}^{2+}$ , $\text{Ba}^{2+}$ , $\text{Pb}^{2+}$ , $\text{Ag}^+$ , or $\text{Ca}^{2+}$ , the resulting compound is insoluble.
Compounds Containing the Following Ions Are Generally Insoluble	Exceptions
$\text{OH}^-$ and $\text{S}^{2-}$	When these ions pair with $\text{Li}^+$ , $\text{Na}^+$ , $\text{K}^+$ , or $\text{NH}_4^+$ , the resulting compounds are soluble.
	When $\text{S}^{2-}$ pairs with $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , or $\text{Ba}^{2+}$ , the resulting compound is soluble.
	When $\text{OH}^-$ pairs with $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , or $\text{Ba}^{2+}$ , the resulting compound is slightly soluble.
$\text{CO}_3^{2-}$ and $\text{PO}_4^{3-}$	When these ions pair with $\text{Li}^+$ , $\text{Na}^+$ , $\text{K}^+$ , or $\text{NH}_4^+$ , the resulting compounds are soluble.

# Model Rules Of Solubility In Aqueous Solutions

**Juan H. Vera, Grazyna Wilczek-Vera**



## **Model Rules Of Solubility In Aqueous Solutions:**

**Ionic Surfactants and Aqueous Solutions** Juan H. Vera, Grazyna Wilczek-Vera, 2018-07-09 Ionic Surfactants and Aqueous Solutions Biomolecules Metals and Nanoparticles covers a wide range of subjects related to aqueous systems from reverse micelles as ion exchangers to the study of micellar phase transfer catalysis for nucleophilic substitution reactions The diverse background expertise and professional interests of the contributors to this book give to it a unique richness of approach in topics of relevance for biotechnology and environmental studies Over sixty publications presenting research results are combined and expanded in this book by some of the original researchers At a mature age and at the summit of successful professional careers they have taken a second look to the state of the art in the fields that they had pioneered Eva Rodil and Ana Soto who had their research formation in the group of Professor Alberto Arce at Universidade de Santiago de Compostela Spain are presently professors at that university Maen Husein is a professor at University of Calgary Canada Remy Dumortier Mohammad Khoshkbarchi Hamid Rabie and Younok Dumortier Shin are presently active leaders in the industrial world in Canada and the USA The editors are retired academics from McGill University Montreal Canada and coauthors of the book **Classical Thermodynamics of Fluid Systems** **Gas Capture Processes** Zhien Zhang, Tohid N. Borhani, Muftah H. El-Naas, Salman Masoudi Soltani, Yunfei Yan, 2020-12-14 This book introduces the recent technologies introduced for gases capture including CO<sub>2</sub> CO SO<sub>2</sub> H<sub>2</sub>S NO<sub>x</sub> and H<sub>2</sub> Various processes and theories for gas capture and removal are presented The book provides a useful source of information for engineers and specialists as well as for undergraduate and postgraduate students in the fields of environmental and chemical science and engineering **Journal of the American Chemical Society** American Chemical Society, 1921 Proceedings of the Society are included in v 1 59 1879 1937 **Elementary principles of chemistry** Raymond Bedell Brownlee, 1921 **Journal of Physical & Colloid Chemistry**, 1928 Includes section New Books *The Journal of Physical Chemistry*, 1925 *Chemical News and Journal of Physical Science*, 1895 *Chemical News and Journal of Industrial Science*, 1887 **Annual Reports on the Progress of Chemistry** Chemical Society (Great Britain), 1908 **Industrial Waste Air Model (IWAIR)**, 2002 *The Chemical News : and Journal of Physical Science*, 1895 **Introduction to Physical Chemistry** Sir James Walker, 1927 *Bombay University Handbook* University of Bombay, 1928 **The Chemical News and Journal of Industrial Science** William Crookes, James H. Gardiner, Gerald Druce, H. W. Blood-Ryan, 1895 **Theoretical Chemistry from the Standpoint of Avogadro's Rule and Thermodynamics** Walther Nernst, 1904 *The Phase Rule and Its Applications* Alexander Findlay, 1923 **Scientific American**, 1909 **Science Abstracts**, 1904 **A Text-book of Inorganic Chemistry for University Students** James Riddick Partington, 1929 *British Chemical Abstracts*, 1926

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