

Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation
- Identify the given (with units) and what you want to find (with units)
- Show set up with units. Check sig figs, give final answer with units and label.

- Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



Before **2 mol** **0 mol** **0 mol**

Change **-2 mol** **+2 mol** **+1 mol**

After **0 mol** **2 mol** **1 mol**

$$36 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.0 \text{ g H}_2\text{O}} = 2.0 \text{ mol H}_2\text{O} \times \frac{2 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} = 2 \text{ mol H}_2 \times \frac{2.0 \text{ g H}_2}{1 \text{ mol H}_2} = 4.0 \text{ g H}_2$$

- Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



Before **9.94 mol** **xs mol** **0 mol**

Change **-9.94 mol** **-0.5 mol** **+9.94 mol**

After **0 mol** **0 mol** **9.94 mol**

$$581 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.5 \text{ g NaCl}} = 9.93 \text{ mol NaCl} \times \frac{2 \text{ mol Na}}{2 \text{ mol NaCl}} = 9.93 \text{ mol Na} \times \frac{23.0 \text{ g Na}}{1 \text{ mol Na}} = 228 \text{ g Na}$$

- You eat 180.0 g of glucose (90 M&Ms). If glucose, $\text{C}_6\text{H}_{12}\text{O}_6$, reacts with oxygen gas to produce carbon dioxide and water, how many grams of oxygen will you have to

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Table of Contents Modeling Chemistry Stoichiometry 2 Percent Yield Answers

1. Understanding the eBook Modeling Chemistry Stoichiometry 2 Percent Yield Answers
 - The Rise of Digital Reading Modeling Chemistry Stoichiometry 2 Percent Yield Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Stoichiometry 2 Percent Yield Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry Stoichiometry 2 Percent Yield Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Stoichiometry 2 Percent Yield Answers
 - Personalized Recommendations
 - Modeling Chemistry Stoichiometry 2 Percent Yield Answers User Reviews and Ratings
 - Modeling Chemistry Stoichiometry 2 Percent Yield Answers and Bestseller Lists

5. Accessing Modeling Chemistry Stoichiometry 2 Percent Yield Answers Free and Paid eBooks
 - Modeling Chemistry Stoichiometry 2 Percent Yield Answers Public Domain eBooks
 - Modeling Chemistry Stoichiometry 2 Percent Yield Answers eBook Subscription Services
 - Modeling Chemistry Stoichiometry 2 Percent Yield Answers Budget-Friendly Options
6. Navigating Modeling Chemistry Stoichiometry 2 Percent Yield Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Modeling Chemistry Stoichiometry 2 Percent Yield Answers Compatibility with Devices
 - Modeling Chemistry Stoichiometry 2 Percent Yield Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry Stoichiometry 2 Percent Yield Answers
 - Highlighting and Note-Taking Modeling Chemistry Stoichiometry 2 Percent Yield Answers
 - Interactive Elements Modeling Chemistry Stoichiometry 2 Percent Yield Answers
8. Staying Engaged with Modeling Chemistry Stoichiometry 2 Percent Yield Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry Stoichiometry 2 Percent Yield Answers
9. Balancing eBooks and Physical Books Modeling Chemistry Stoichiometry 2 Percent Yield Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry Stoichiometry 2 Percent Yield Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modeling Chemistry Stoichiometry 2 Percent Yield Answers
 - Setting Reading Goals Modeling Chemistry Stoichiometry 2 Percent Yield Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modeling Chemistry Stoichiometry 2 Percent Yield Answers
 - Fact-Checking eBook Content of Modeling Chemistry Stoichiometry 2 Percent Yield Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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