At a high temperature, 0.50 mol of HBr was placed in a 1.0 L container and allowed to decompose according to the reaction:

$$2HBr(g) \rightleftharpoons H_{2(g)} + Br_{2(g)}$$

At equilibrium the [Br<sub>2</sub>] was measured to be 0.13 M. What is K<sub>eq</sub> for this reaction at this temperature?

	2HBr =	= H2 +	Br2
[x]	0.50	0	0
[C]	-0-26	+0.13	+0.13
IE]	0.24	0.13	0.13

Keg =	IH2][Br]2	= (0.13)2	=0.29
	100001100000000000000000000000000000000	150000000000000	

Answer Keg = 0.29

 When 1.0 mol of NH<sub>3(g)</sub> and 0.40 mol of N<sub>2(g)</sub> are placed in a 5.0 L vessel and allowed to reach equilibrium at a certain temperature, it is found that 0.78 mol of NH<sub>3</sub> is present. The reaction is:

2NH3(g) = 3H2(g) + N2(g)

Initial [NH3] = 1.0 mol = 0.20 M Initial [N2] = 0.40 mol = 0.080 M

Equil [NH3] = 0.78 mol = 0.156 M 2NH2 = 3 H2 + N2

	aNH3 &	= 3 Na	+ N2
[I]	0.20	0	0.080
EC3	-0.044	+0.066	+0.022
LEJ.	0.156	0.066	0.102

a) Calculate the equilibrium concentrations of all three species.

[NH<sub>3</sub>] - 0.16 M [H<sub>2</sub>] - 0.066 M [N<sub>2</sub>] - 0.10 M

b), Calculate the value of the equilibrium constant at this temperature. (use unrounded cones, then round to 2503)

 $\text{Keq} = \frac{\text{EH}_2 \text{J}^3 \text{EN}_2 \text{J}}{\text{ENH}_3 \text{J}^2} = \frac{(0.066)^3 (0.102)}{(0.156)^2} = 0.0012$ Answ

Answer 1.2×10-3

c) How many moles of H<sub>2</sub> are present at equilibrium?

0.066 M x 5.0 L = 0.33 mal

Answer 0.33 mol

d) How many moles of N<sub>2</sub> are present at equilibrium?

0.102 M x 5.0 L = 0.51 mal

Answer 0.51 mol

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# Table of Contents Modeling Chemistry U6 Ws 3 V2 Answers

- 1. Understanding the eBook Modeling Chemistry U6 Ws 3 V2 Answers
  - The Rise of Digital Reading Modeling Chemistry U6 Ws 3 V2 Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Modeling Chemistry U6 Ws 3 V2 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Chemistry U6 Ws 3 V2 Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Modeling Chemistry U6 Ws 3 V2 Answers
  - Personalized Recommendations
  - Modeling Chemistry U6 Ws 3 V2 Answers User Reviews and Ratings
  - Modeling Chemistry U6 Ws 3 V2 Answers and Bestseller Lists

- 5. Accessing Modeling Chemistry U6 Ws 3 V2 Answers Free and Paid eBooks
  - o Modeling Chemistry U6 Ws 3 V2 Answers Public Domain eBooks
  - Modeling Chemistry U6 Ws 3 V2 Answers eBook Subscription Services
  - Modeling Chemistry U6 Ws 3 V2 Answers Budget-Friendly Options
- 6. Navigating Modeling Chemistry U6 Ws 3 V2 Answers eBook Formats
  - ∘ ePub, PDF, MOBI, and More
  - Modeling Chemistry U6 Ws 3 V2 Answers Compatibility with Devices
  - Modeling Chemistry U6 Ws 3 V2 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modeling Chemistry U6 Ws 3 V2 Answers
  - Highlighting and Note-Taking Modeling Chemistry U6 Ws 3 V2 Answers
  - Interactive Elements Modeling Chemistry U6 Ws 3 V2 Answers
- 8. Staying Engaged with Modeling Chemistry U6 Ws 3 V2 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - $\circ\,$  Following Authors and Publishers Modeling Chemistry U6 Ws 3 V2 Answers
- 9. Balancing eBooks and Physical Books Modeling Chemistry U6 Ws 3 V2 Answers
  - $\circ$  Benefits of a Digital Library
  - $\circ\,$  Creating a Diverse Reading Collection Modeling Chemistry U6 Ws 3 V2 Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry U6 Ws 3 V2 Answers
  - Setting Reading Goals Modeling Chemistry U6 Ws 3 V2 Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry U6 Ws 3 V2 Answers
  - Fact-Checking eBook Content of Modeling Chemistry U6 Ws 3 V2 Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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