

7 • Molecular Structure

STATION 3 • POLARITY

Determine the shape of the molecule and decide whether each compound is polar or nonpolar.

PH_3 $3+1=8$ $\text{H} : \underset{\cdot\cdot}{\underset{\cdot\cdot}{\text{P}}} : \text{H}$ Polar	CF_4 $2+4=8$ $\begin{array}{c} :\ddot{\text{F}}: \\ :\ddot{\text{F}}: \text{C} : \ddot{\text{F}}: \\ :\ddot{\text{F}}: \end{array}$ non-polar	BH_3 $3+3=6$ $\text{H} : \underset{\cdot\cdot}{\underset{\cdot\cdot}{\text{B}}} : \text{H}$ non-polar
SO_2 $6+6=12$ $\begin{array}{c} :\ddot{\text{O}}: \text{S} : \ddot{\text{O}}: \\ \text{or } :\ddot{\text{O}}: \text{S} : \ddot{\text{O}}: \\ \quad \quad \quad :\ddot{\text{O}}: \text{S} : \ddot{\text{O}}: \end{array}$ Polar	CO_2 $12+4=16$ $\begin{array}{c} :\ddot{\text{O}}: \text{C} : \ddot{\text{O}}: \\ \quad \quad \quad :\ddot{\text{O}}: \text{C} : \ddot{\text{O}}: \end{array}$ non-polar	SCl_2 $6+6=12$ $\begin{array}{c} :\ddot{\text{Cl}}: \text{S} : \ddot{\text{Cl}}: \\ \quad \quad \quad :\ddot{\text{Cl}}: \text{S} : \ddot{\text{Cl}}: \end{array}$ polar

7 • Molecular Structure

STATION 4 • IMF's AND ATTRACTIONS

Match the statement with the Inter-Particle Forces

A. Dipole-Dipole Interactions

B. Covalent bonds

C. Hydrogen-bonding

D. Ionic Bonds

E. Metallic Bonds

F. London Dispersion Forces

- F 1. Is used to explain why BP of $\text{Xe} > \text{Kr} > \text{Ar} > \text{Ne} > \text{He}$
- B 2. Is present in ~~Complete~~ all ~~Compounds~~ Compounds
- E 3. Is used to explain why Cu is a good conductor
- D 4. Is used to explain why NaCl(s) is a good conductor
- A 5. Is used to explain why ICI has a higher BP than Br_2
- C 6. Is used to explain why boiling point of CH_3OH is higher than CH_4
- E 7. A lattice of positive ions in a sea of delocalized electrons
- A 8. Positive ends of molecules attract negative ends of other molecules
- F 9. Lattice of alternating positively and negatively charged particles
- C 10. Hydrogen atoms tethered between small, highly electronegative atoms N, O, and F.
- F 11. Polarized electron clouds induce dipoles in their neighboring atoms

Modeling Chemistry Unit 7 Review Answers

ML Yell



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