

Chemistry Unit 8 Worksheet 4

Examples of Every Kind of Problem

This is a separate sheet of paper, so this is a complete solution to each of the problems. You can follow the problems without an answer. We want to make your final solution.

12. Calculate the number of moles of potassium chlorate, KClO_3 , that must decompose to produce potassium chloride KCl and a total of 1.5 moles of oxygen gas.



| | | | |
|---------|------|------|------|
| Subst: | 1.0 | 0 | 0 |
| Change: | -1.0 | +1.0 | +1.5 |
| Final: | 0 | 1.0 | 1.5 |

1.0 moles KClO_3 needed

13. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 3.2 g of magnesium?



| | | | | |
|---------|--------|--------|--------|--------|
| Subst: | 0.100 | 0 | 0 | 0 |
| Change: | -0.100 | -0.200 | +0.100 | +0.100 |
| Final: | 0 | 0 | 0.100 | 0.100 |

0.200 moles HCl needed

14. Ethane, C_2H_6 , reacts with oxygen gas to produce carbon dioxide gas and water vapor. What number of moles of oxygen are required to react with 1.0 mole of ethane?



| | | | | |
|---------|--------|--------|--------|--------|
| Subst: | 1.000 | 0.00 | 0 | 0 |
| Change: | -1.000 | -3.500 | +2.000 | +3.000 |
| Final: | 0 | 3.50 | 2.000 | 3.000 |

3.5 moles O_2 needed

15. Sodium metal reacts with sulfuric acid to produce sodium sulfate and hydrogen gas. How many moles of sulfuric acid are needed to react with 1.0 mole of sodium metal?



| | | | | |
|---------|---------|---------|---------|---------|
| Subst: | 1.0000 | 0.00 | 0 | 0 |
| Change: | -1.0000 | -0.5000 | +0.5000 | +1.0000 |
| Final: | 0 | 0.50 | 0.5000 | 1.0000 |

0.5000 moles H_2SO_4 needed

Modeling Chemistry Unit 8 4

SA Adler



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