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Passage 1 (Questions 1-4)

To test how amino acids first appeared on this planet, an experiment was set up to simulate the conditions thought to have existed on early Earth. H_2O , CH_4 , NH_3 , and H_2 were all introduced into a sterile arrangement of sealed flasks and tubes. Liquid water evaporated in one compartment as "lightning" was induced between a pair of electrodes. The water vapor then cooled and condensed in a connected section, which fed back to the first compartment, until a dynamic equilibrium was reached.

After two weeks of continuous operation, roughly 2% of the carbon in the system was in the form of amino acids, with glycine being the most abundant. The following steps are believed to account for the formation of the key intermediates, formaldehyde and hydrogen cyanide.



Mechanism 1

Via the Strecker Synthesis, glycine can then form as follows:



Mechanism 2

Several other amino acids were also observed to form. Of the chiral amino acids, both enantiomers were created in a racemic mixture, though in most biological systems L-amino acids are used exclusively for protein synthesis.

Simple sugars also formed in the same environment. Once formaldehyde is present to react with H_2O , it does so according to the Butlerov reaction:



Figure 1. Formation of carbohydrates from aqueous formaldehyde

Question 4

Which of the following will increase the yield of carboxylic acid formed from a nitrile group-containing compound, using a mechanism similar to the Strecker synthesis?

- ☐ A. Use a stronger nucleophile, like a Grignard Reagent, instead of H_2O
- ☐ B. Attach an electron withdrawing group to the carbon of the nitrile group
- ☐ C. Use a stronger electrophile, like a Grignard Reagent, instead of H_2O
- ☐ D. Attach an electron donating group to the carbon of the nitrile group

Kaplan Mcat Practice Test 1 Answers

Michael Hoelscher



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Table of Contents Kaplan Mcat Practice Test 1 Answers

1. Understanding the eBook Kaplan Mcat Practice Test 1 Answers
 - The Rise of Digital Reading Kaplan Mcat Practice Test 1 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Kaplan Mcat Practice Test 1 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Kaplan Mcat Practice Test 1 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Kaplan Mcat Practice Test 1 Answers
 - Personalized Recommendations
 - Kaplan Mcat Practice Test 1 Answers User Reviews and Ratings
 - Kaplan Mcat Practice Test 1 Answers and Bestseller Lists
5. Accessing Kaplan Mcat Practice Test 1 Answers Free and Paid eBooks
 - Kaplan Mcat Practice Test 1 Answers Public Domain eBooks
 - Kaplan Mcat Practice Test 1 Answers eBook Subscription Services
 - Kaplan Mcat Practice Test 1 Answers Budget-Friendly Options

6. Navigating Kaplan Mcat Practice Test 1 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Kaplan Mcat Practice Test 1 Answers Compatibility with Devices
 - Kaplan Mcat Practice Test 1 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Kaplan Mcat Practice Test 1 Answers
 - Highlighting and Note-Taking Kaplan Mcat Practice Test 1 Answers
 - Interactive Elements Kaplan Mcat Practice Test 1 Answers
8. Staying Engaged with Kaplan Mcat Practice Test 1 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Kaplan Mcat Practice Test 1 Answers
9. Balancing eBooks and Physical Books Kaplan Mcat Practice Test 1 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Kaplan Mcat Practice Test 1 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Kaplan Mcat Practice Test 1 Answers
 - Setting Reading Goals Kaplan Mcat Practice Test 1 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Kaplan Mcat Practice Test 1 Answers
 - Fact-Checking eBook Content of Kaplan Mcat Practice Test 1 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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