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$$m_{\text{mix}} = \frac{m_1 + m_2}{V_1 + V_2} \times \frac{1}{1000} \times \frac{1}{M}$$
 where  $m_1$  and  $m_2$  are masses in grams,  $V_1$  and  $V_2$  are volumes in ml, and  $M$  is molar mass in g/mol. For example, if  $m_1 = 480$  g,  $m_2 = 520$  g,  $V_1 = 720$  ml,  $V_2 = 624$  ml, and  $M = 1000$  g/mol, then  $m_{\text{mix}} = \frac{480 + 520}{720 + 624} \times \frac{1}{1000} \times \frac{1}{1} = 0.001$  mol/L. With increase in temperature which of these changes a molality b fraction of solute present in water

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