



Le Jeu De Larmurier

Wlodimir Konarski



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mol ko 2 and 0 mol h₂ o what is the limiting reactant

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$$4\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu} + 2\text{Ag} + 2\text{Cu(NO}_3)_2$$

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$$3\text{Co} + 7\text{H}_2\text{C}_3\text{H}_8 \rightarrow 3\text{H}_2\text{C}_3\text{H}_8 + 2\text{O}$$

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web start with 8 3 mole h₂s and multiply it with 2 mole h₂o over 2 mole h₂s from the equation

$$8 3 \text{ mole h}_2\text{s} \times \frac{2 \text{ mole h}_2\text{o}}{2 \text{ mole h}_2\text{s}} = 8 3 \text{ mole h}_2\text{o}$$

 8 3 mole h₂s x 18g h₂o i mole h₂o 149 4g h₂o is your theoretical yield so percent yield 137 1g h₂o actual yield 149 4g h₂o x 100
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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

 if the reaction of 21 3 grams of o₂ produces a 47 3 yield how many grams of co would be produced step one figure out how many moles of the limiting reagent you have

$$\frac{21 3 \text{ grams}}{32 \text{ g/mol}} = 0 66 \text{ moles}$$

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 calculated mass of the product expected x100 in this reaction a yield 80 is anticipated

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80 g I_2O_5 1 mol I_2O_5 1 mol I_2 xs 1 333.8 g I_2O_5 1 mol I_2O_5 28 g CO 1 mol CO

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