



Mercedes Benz B200 Owner Manual

H Kauffman



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monoxide co determine the mass of iodine I_2 which could be produced b if in the above situation only 0.160 moles of iodine I_2 was produced i what mass of iodine was produced

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$$Cu + 2AgNO_3 \rightarrow Cu(NO_3)_2 + 2Ag$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } H_2S \times \frac{2 \text{ mole } H_2O}{2 \text{ mole } H_2S} = 8.3 \text{ mole } H_2O$ $8.3 \text{ mole } H_2O \times 18 \text{ g } H_2O = 149.4 \text{ g } H_2O$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } H_2O}{149.4 \text{ g } H_2O} \times 100 = 91.77\%$ problems 6 8 are similar to problem 5

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you have 0.216232 grams mole 21.3 grams 0.66 moles 32 g m

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calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ 4 $\text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0

g $\text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}_2\text{O}_3$ 4 mol $\text{Al} \times 1 \text{ mol Al}_2\text{O}_3$ c determine the percent yield

yield actual $\times 100$ 260.0 g $\times 100$ 86.01 theoretical 302.3 g

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web 7 Na_2CO_3 is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be

weighed by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12

C is the limiting reagent 4 33 g of H_2 are left over 13

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from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s

remains when the reaction is complete

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web the percent yield is calculated as follows $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent

yield theoretical $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ based on this definition we would expect a percent yield

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CH_3OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5 g of CO is

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reaction

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