

NAME: ADM NO: CLASS:
233/3
CHEMISTRY
PRACTICAL
FORM THREE
END OF TERM 1 EXAMS -2023
TIME: 2 HOURS

You are provided with the following:

- 3.3g metal carbonate, MCO_3 , labeled solution Q
- 2M hydrochloric acid, labeled solution P
- Sodium hydroxide, labeled solution R containing 40g/L of solution

You are required to determine the relative atomic mass of metal M

Procedure

- Measure accurately 100cm^3 of solution P into clean 250cm^3 conical flask and add all the 3.3g of solid Q, MCO_3
- Shake the mixture well and wait for effervescence to stop. Label the resulting solution as S
- Pipette 25cm^3 of solution R into a conical flask and add 2-3 drops of phenolphthalein indicator.
- Fill the burette with solution S and titrate against the solution R until the end point.
- Record your results in the table below. Repeat the procedure at least two times to complete the table.
(4 mks)

	i	ii	iii
Final burette reading (cm^3)			
Initial burette reading (cm^3)			
Volume of solution S used (cm^3)			

- What is the average volume of solution S used? (1mk)
- Calculate the moles of sodium hydroxide, solution R used. (2mks)
- Calculate the moles of hydrochloric acid in the average volume of solution S used. (2mks)

June 2013 Chemistry Paper 3

Camilla Rothe

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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