

Sodium becomes a cation, and chlorine becomes an anion.

|                                       |   |   |
|---------------------------------------|---|---|
| <b>Activity A:</b><br><br><b>Ions</b> | <b>Get the Gizmo ready:</b> <ul style="list-style-type: none"> <li>Click <b>Reset</b>.</li> <li>Check that sodium and chlorine are still selected.</li> </ul> |  |
|---------------------------------------|---|---|

**Introduction:** Some of the particles that make up atoms have an electrical charge. Electrons are negatively charged, while protons are positively charged. Particles with opposite charges (+ and -) attract, while particles with the same charge (+ and + or - and -) repel.

**Question:** What happens when atoms gain or lose electrons?

1. **Count:** Electrons move around the nucleus of atoms in specific shells, shown by the rings around the atoms in the Gizmo. The first ring holds two electrons, and the second holds eight. (Electrons in the inner rings are not shown; you can assume these rings are full.)

- A. Observe the sodium and chlorine atoms. Assuming that the inner rings are full of electrons, how many electrons are there total in each atom?

Sodium: 11 electrons

Chlorine: 17 electrons

- B. Each atom is neutrally charged, which means that each atom has the same number of protons and electrons. Based on this, how many protons are in each atom?

Sodium: 11 protons

Chlorine: 17 protons

2. **Observe:** Most atoms are stable with a configuration of eight valence electrons. This is known as the **octet rule**. How many valence electrons does each atom have?

Sodium: 1 valence electron

Chlorine: 7 valence electrons

3. **Form a bond:** Each electron has a charge of  $1-$ , and each proton has a charge of  $1+$ . You can calculate the charge of an atom by subtracting the number of electrons from the number of protons. Move an electron from the sodium to the chlorine atom.

- A. What are the charges of each atom now? Sodium: +1 Chlorine: -1

Turn on **Show charge** to check. These charged atoms are called **ions**.

- B. Is each ion stable? Explain.

Yes, each ion is stable. Chlorine only has 7 valence electrons, so gaining 1 more would lead to a stable full octet. Sodium on the other hand has 1 valence electron, by losing this electron it has just a regular shell of 8 electrons, also becoming a stable octet.

Click **Check** in the lower right corner of the Gizmo to check.

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# Ionic Bonds Gizmo Activity A Ions Key

**S Nieto**



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