

CK-12 Chemistry - Basic
Answer Key**Chapter 8: Ionic and Metallic Bonding**

8.1 Ions**Practice 8.1**

Write the ground state configuration for the metal calcium, and predict the ion form it must become to be isoelectric with the nearest noble gas.

Answer:

Ca: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$ or $[\text{Ar}]4s^2$

The nearest noble gas to calcium is $[\text{Ar}]$. So calcium will typically lose two electrons to become isoelectric with $[\text{Ar}]$.

Practice 8.2

Write the ground state configuration for the nonmetal sulfur, and predict the ion form it must become to be isoelectric with the nearest noble gas.

Answer:

S: $1s^2 2s^2 2p^6 3s^2 3p^4$ or $[\text{Ne}]3s^2 3p^4$

The nearest noble gas to sulfur is $[\text{Ar}]$, so sulfur will typically gain two electrons to become $[\text{Ne}] 3s^2 3p^6$ which is isoelectric with $[\text{Ar}]$

Lesson Review Questions*Questions*

1. Draw electron dot diagrams for one metal and one nonmetal from the periodic table.
2. Predict which of the following will likely become a cation or an anion.
 - a. Ca
 - b. Na
 - c. F
 - d. Br
 - e. S
3. Write the ground state configuration for the following elements and predict the ion form it must become to be isoelectric with the nearest noble gas.
 - a. Be

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