



CHEMICAL BONDS



IONIC, COVALENT, & METALLIC

FORMS OF CHEMICAL BONDS
Bonding: combination of atoms

Atoms combine to form molecules or compounds by sharing or transferring electrons. This process is called bonding. The resulting molecule or compound has properties different from the individual atoms.

Types of Bonds:

- Covalent Bond:** Formed by sharing electrons between two atoms. Example: H_2 , H_2O , CO_2 .
- Ionic Bond:** Formed by the transfer of electrons from one atom to another, resulting in oppositely charged ions that attract each other. Example: $NaCl$, MgO .
- Metallic Bond:** Formed by the attraction between positive metal ions and a sea of delocalized electrons. Example: Fe , Cu , Al .

Properties of Bonds:

- Covalent bonds are strong and directional.
- Ionic bonds are strong but non-directional.
- Metallic bonds are strong and non-directional.

Science Real

A simplified periodic table of elements, color-coded by groups.

Diagram of Ionic Bonding: Shows a red circle with a minus sign ($-$) and a yellow circle with a plus sign ($+$) connected by a dashed line, representing an ionic bond.

Diagram of Covalent Bonding: Shows two atoms, each with a nucleus and electron shells. A red arrow points to the overlapping region of the outer shells, where electrons are shared to form a covalent bond.

Diagram of Metallic Bonding: Shows a cluster of positive metal ions (yellow circles with plus signs) surrounded by a sea of delocalized electrons (blue dots).

Text on the right:

When two atoms share a pair of electrons, a covalent bond is formed. The resulting molecule has properties different from the individual atoms.

Covalent bonds are strong and directional. They are formed by the sharing of electrons between two atoms.

Ionic bonds are strong but non-directional. They are formed by the transfer of electrons from one atom to another, resulting in oppositely charged ions that attract each other.

Metallic bonds are strong and non-directional. They are formed by the attraction between positive metal ions and a sea of delocalized electrons.

GUIDED READING ACTIVITY

Ionic Bonds Chapter 5 Reading Notetaking Guide Answers

Terry C. Jones



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