

Name: Sam Li

Simulation: Ionic and Covalent Bonding

Background

In this investigation you will bond select atoms. Based upon the types of atoms that you choose to combine, you will create either an ionic compound or a covalent compound. You will have the opportunity to analyze the differences between these different types of compounds and to predict the number of atoms needed to create each, as well as learn how to appropriately name them.

1. Describe the difference between an atom and a molecule:

Atoms are single particles, and molecules are made up of multiple atoms

2. Where are metal atoms located on the periodic table? Where are non-metal atoms located on the periodic table?

Metals are on the left side and non-metals are on the right

3. What subatomic particle(s) participate in chemical bonding?

Electron

4. In your own words, define *valence electron*.

electrons on outside ring which allow for bonding, negative charge

5. How can you determine the number of valence electrons in an atom using the periodic table?

Look at the atomic #

6. Draw a Lewis Dot Structure for the following atoms:

a. Strontium (Sr)



b. Carbon (C)



c. Iodine (I)



d. Xenon (Xe)

*Check your answers before moving on to the next portion of the activity.

Ionic Or Covalent Bonding Lab Answers

Annelies Wilder-Smith



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Study Guide and Selected Solutions Karen C. Timberlake, 2001-11 Keyed to the learning goals in the text this guide is designed to promote active learning through a variety of exercises with answers and mastery exams The guide also contains complete solutions to odd numbered problems **The Essential Lab Manual** Karen Timberlake, 2002-06-24 Drawing from the successful main Laboratory Manual the Essential Laboratory Manual includes twenty one experiments which have been revised and updated Suitable for a one or two term lab course Prentice Hall Science Explorer Michael J. Padilla, 2002

Chemical Interactions, 2005 **Chemistry** Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02 *Polarity, Solutions, and Separation Science* Kenda Jo Lemont, 1998 **A Laboratory Centered Molecular Biology Teaching Module which Facilitates the Teaching of Biological Science** Michael James Brundage, 1990 **Using Computers in Chemistry and Chemical Education** Theresa Julia Zielinski, Mary L. Swift, 1997 Based on how computers are used in research and industry this timely volume provides a practical curriculum for using computers in training chemists and other professionals It spans the full range of applications from spreadsheets to specialized software for ab initio calculations With contributions from experts in a variety of fields the book will be invaluable for anyone developing a college level course in chemistry *Cooperative Chemistry Lab Manual* Cooper, 2005-02 The laboratory course described in the lab manual emphasizes experimental design data analysis and problem solving Inherent in

the design is the emphasis on communication skills both written and oral Students work in groups on open ended projects in which they are given an initial scenario and then asked to investigate a problem There are no formalized instructions and students must plan and carry out their own investigations *Glencoe Science* McGraw-Hill Staff,2001-06 **Chemistry in the Community** ,1995 **Focus on physical science** Charles H. Heimler,1989 **Catalog Issue** University of Colorado,1962 **Introduction to Physical Science** James T. Shipman,2000-06

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