

- Directions:**
 1. An element may be assigned a positive or negative oxidation number.
 2. An atom with a higher electronegativity is assigned a negative oxidation number.
 3. The sum of all oxidation numbers for a molecule must be zero.

Ion Practice Set

1. Write an equation that shows how the element is oxidized or reduced.
 2. Write the oxidation number next to the element.
 3. Write the oxidation number next to the element.

Complete the following table, using the periodic table in the back of your book.

| | ELEMENT NAME | ION SYMBOL | NUMBER OF PROTONS | NUMBER OF ELECTRONS | NUMBER OF ELECTRONS LOST OR GAINED | TYPE OF ION |
|----|--------------|------------|-------------------|---------------------|------------------------------------|-------------|
| 1 | Sodium | Na^+ | 11 | 10 | lost one | cation |
| 2 | Chlorine | Cl^- | 17 | 18 | gained one | anion |
| 3 | Sulfur | S^{2-} | 16 | 18 | gained two | anion |
| 4 | Indium | In^{3+} | 49 | 46 | lost three | cation |
| 5 | Calcium | Ca^{2+} | 20 | 18 | lost two | cation |
| 6 | Strontium | Sr^{2+} | 38 | 36 | lost two | cation |
| 7 | Hydrogen | H^+ | 1 | 0 | lost one | cation |
| 8 | Lithium | Li^+ | 3 | 2 | lost one | cation |
| 9 | Aluminum | Al^{3+} | 13 | 10 | lost three | cation |
| 10 | Barium | Ba^{2+} | 56 | 54 | lost two | cation |
| 11 | Selenium | Se^{2-} | 34 | 36 | gained two | anion |
| 12 | Hydrogen | H^+ | 1 | 0 | lost one | cation |
| 13 | Strontium | Sr^{2+} | 38 | 36 | lost two | cation |
| 14 | Barium | Ba^{2+} | 56 | 54 | lost two | cation |
| 15 | Calcium | Ca^{2+} | 20 | 18 | lost two | cation |
| 16 | Indium | In^{3+} | 49 | 46 | lost three | cation |
| 17 | Chlorine | Cl^- | 17 | 18 | gained one | anion |

Ion And Isotope Practice Set Answer Key

Judd E. Hollander



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