

## Ionic Equations

Ionic equations show us which **species** (atoms, ions and molecules) are involved in a reaction.

An ionic substance splits up into its ions.

E.g.  $\text{NaCl}$  forms  $\text{Na}^+ + \text{Cl}^-$

This also happens if an acid or soluble base dissolves in water.

Must be (aq) – not (s) or (g)

E.g.  $\text{HCl (aq)}$  forms  $\text{H}^+ + \text{Cl}^-$      $\text{Mg(OH)}_2 \text{ (aq)}$  forms  $\text{Mg}^{2+} + 2\text{OH}^-$     but  $\text{Ca(OH)}_2 \text{ (s)}$  does not dissociate

Pure elements and covalent substances (except acids) will not form ions.

E.g.  $\text{Mg}$ ,  $\text{C}$ ,  $\text{CO}_2$ ,  $\text{H}_2$

1. Circle the substances that will NOT form ions:

$\text{NaCl}$      $\text{CO}_2$      $\text{CuCO}_3$      $\text{H}_2$      $\text{SiO}_2$      $\text{Al}$      $\text{C}_2\text{H}_6$      $\text{Mg(OH)}_2$      $\text{LiCl}$      $\text{PbI}_2$      $\text{H}_2\text{O}$

2. Identify the species in these substances:

a)  $\text{MgSO}_4$      $\text{Mg}^{2+} + \text{SO}_4^{2-}$

b)  $\text{CaCl}_2$

c)  $\text{Li}_2\text{O}$

d)  $\text{HNO}_3$

e)  $2\text{Na}_2\text{O}$

f)  $2\text{CaCO}_3$

g)  $\text{H}_2\text{SO}_4$  (sulfuric acid)

h)  $\text{HNO}_3$  (nitric acid)

i)  $\text{Al}_2\text{O}_3$

j)  $\text{MgCO}_3$

k)  $\text{Fe(OH)}_3$

l)  $\text{Na}_2\text{SO}_4$

3. Identify the species in these substances:

a)  $\text{Mg} + 2\text{HCl}$      $\text{Mg} + 2\text{H}^+ + 2\text{Cl}^-$

b)  $\text{MgCl}_2 + \text{H}_2$

c)  $\text{Li} + \text{O}_2$

d)  $\text{CaCO}_3 + 2\text{HCl}$

e)  $2\text{Na}_2\text{O} + 2\text{Li}$

f)  $2\text{Zn} + \text{H}_2\text{SO}_4$

g)  $\text{MgO} + \text{CO}_2$

h)  $2\text{NaI} + \text{Br}_2$

i)  $\text{Ca} + 2\text{HNO}_3$

4. Identify the species in the reactants and products:

a)  $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$   
 $\text{Mg} + 2\text{H}^+ + 2\text{Cl}^- \rightarrow \text{Mg}^{2+} + 2\text{Cl}^- + \text{H}_2$

b)  $4\text{Li} + \text{O}_2 \rightarrow 2\text{Li}_2\text{O}$

c)  $\text{CaCl}_2 + 2\text{Li} \rightarrow 2\text{LiCl} + \text{Ca}$

d)  $\text{Ca} + \text{H}_2\text{SO}_4 \rightarrow \text{CaSO}_4 + \text{H}_2$

e)  $2\text{Al}_2\text{O}_3 \rightarrow 4\text{Al} + 3\text{O}_2$

f)  $\text{MgCl}_2 + \text{F}_2 \rightarrow \text{MgF}_2 + \text{Cl}_2$

# Ionic Equations Answers Instructional Horizons

**Thomas Griffiths**



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