

Name Solutions

Date _____ Period _____

Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1 ve	transfers $1e^-$ to F	Li^+
F	7 ve	accepts $1e^-$ from Li	F^-

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $Li + Cl \Rightarrow LiCl$	Li	1 ve	donates $1e^-$	Li^+
$Li + Cl \Rightarrow LiCl$	Cl	1 ve	accepts $1e^-$	Cl^-
2) $Ca + O \Rightarrow CaO$	Ca	2 ve	donates $2e^-$	Ca^{+2}
$Ca + O \Rightarrow CaO$	O	6 ve	accepts $2e^-$	O^{2-}
3) $Be + F \Rightarrow BeF_2$	Be	2 ve	donates $2e^-$ to each F	Be^{+2}
$Be + F \Rightarrow BeF_2$	F	7 ve	each F accepts $1e^-$ each	F^-
4) $Mg + S \Rightarrow MgS$	Mg	2 ve	donates $2e^-$ to S	Mg^{+2}
$Mg + S \Rightarrow MgS$	S	6 ve	accepts $2e^-$	S^{2-}
5) $K + F \Rightarrow KF$	K	1 ve	donates $1e^-$	K^+
$K + F \Rightarrow KF$	F	7 ve	accepts $1e^-$	F^-

Ionic Metallic Bonding Guided Answer Key

S Ben Porath



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Table of Contents Ionic Metallic Bonding Guided Answer Key

1. Understanding the eBook Ionic Metallic Bonding Guided Answer Key
 - The Rise of Digital Reading Ionic Metallic Bonding Guided Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Metallic Bonding Guided Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Metallic Bonding Guided Answer Key
 - User-Friendly Interface

4. Exploring eBook Recommendations from Ionic Metallic Bonding Guided Answer Key
 - Personalized Recommendations
 - Ionic Metallic Bonding Guided Answer Key User Reviews and Ratings
 - Ionic Metallic Bonding Guided Answer Key and Bestseller Lists
5. Accessing Ionic Metallic Bonding Guided Answer Key Free and Paid eBooks
 - Ionic Metallic Bonding Guided Answer Key Public Domain eBooks
 - Ionic Metallic Bonding Guided Answer Key eBook Subscription Services
 - Ionic Metallic Bonding Guided Answer Key Budget-Friendly Options
6. Navigating Ionic Metallic Bonding Guided Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Ionic Metallic Bonding Guided Answer Key Compatibility with Devices
 - Ionic Metallic Bonding Guided Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Metallic Bonding Guided Answer Key
 - Highlighting and Note-Taking Ionic Metallic Bonding Guided Answer Key
 - Interactive Elements Ionic Metallic Bonding Guided Answer Key
8. Staying Engaged with Ionic Metallic Bonding Guided Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic Metallic Bonding Guided Answer Key
9. Balancing eBooks and Physical Books Ionic Metallic Bonding Guided Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Metallic Bonding Guided Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ionic Metallic Bonding Guided Answer Key
 - Setting Reading Goals Ionic Metallic Bonding Guided Answer Key
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Ionic Metallic Bonding Guided Answer Key
 - Fact-Checking eBook Content of Ionic Metallic Bonding Guided Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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