

Use the periodic table and your knowledge of periodic trends to answer the following questions.

Atomic Radius

- Which atom in each pair has the larger atomic radius? Circle your answer.
 a) Li or K b) Ca or Ni c) Ga or B d) O or C e) Cl or Br
 f) Be or Ba g) Si or S h) Fe or Au
- Why do atoms get smaller as you move left to right in a period? Larger as you move down a group?
- Arrange the following in order of increasing atomic radius.
 Mg, Na, P, Si and Ar _____

Ionic Radius

- Explain the relationship between the relative size of an ion to its neutral atom and the charge on the ions.
- Which particle has the larger radius in each atom/ion pair?
 a) Na, Na⁺ b) S, S²⁻ c) I, I⁻ d) Al, Al³⁺ e) Cl, Cl⁻

Ionization Energy

- Which element in each pair has the larger ionization energy?
 a) Na or O b) Be or Ba c) Ar or F d) Cu or Ra e) I or Ne
 f) K or V g) Ca or Fr h) W or Se
- Arrange the following groups of elements in order of increasing ionization energy.
 a) Be, Mg, Sr _____ b) Bi, Cs, Ba _____ c) Na, Al, S _____

Electronegativity

- Which element in each pair has a higher electronegativity value?
 a) Cl, F b) C, N c) Mg, Ne d) As, Ca e) Fr, Cs

Metallic Character

- Arrange the elements from least to most metallic: Sodium, barium, cesium, magnesium, calcium

Electron Affinity

- Write the chemical equation for the addition of an electron to a chlorine atom.
- Arrange the following atoms in order of **increasing** electron affinity: Br, Sb, I, Te, Cl

Ionic Radius Answer Key

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