

5 Descriptions of the bonding in three substances are given.

- substance 1 strong covalent bonds between atoms, permanent dipole-dipole attractions between molecules
- substance 2 strong covalent bonds between atoms, weak forces between molecules
- substance 3 strong covalent bonds between atoms, hydrogen bonding between molecules

Which compounds could be substances 1, 2 and 3?

	substance 1	substance 2	substance 3
<b>A</b>	CH <sub>3</sub> OH	Al <sub>2</sub> Cl <sub>6</sub>	CH <sub>2</sub> Cl <sub>2</sub>
<b>B</b>	Al <sub>2</sub> Cl <sub>6</sub>	CH <sub>2</sub> Cl <sub>2</sub>	CH <sub>4</sub>
<b>C</b>	CH <sub>2</sub> Cl <sub>2</sub>	CH <sub>4</sub>	CH <sub>3</sub> OH
<b>D</b>	CH <sub>4</sub>	CH <sub>3</sub> OH	H <sub>2</sub> O

6 Which type of bonding is **never** found in elements?

- A** covalent
- B** ionic
- C** metallic
- D** van der Waals' forces

7 Using the information in the table, what is the enthalpy change,  $q$ , for the reaction described?



standard enthalpy change	value / kJ mol <sup>-1</sup>
$\Delta H_{\text{f}}^\circ \text{ CsBr}(\text{s})$	+25.9
$\Delta H_{\text{f}}^\circ \text{ Cs}^+(\text{g})$	-276
$\Delta H_{\text{f}}^\circ \text{ Br}^-(\text{g})$	-335

- A** -636.9      **B** -585.1      **C** +585.1      **D** +636.9

# May 2013 9701 A Level Chemistry Paper

**J Dewey**



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