

CHEMICAL THERMODYNAMICS

- 1) In open system, system and surroundings exchange.....
 (a) Energy only (b) Matter only
(c) Energy and matter (d) Neither energy nor matter
- 2) An isothermal expansion takes place at.....
 (a) constant pressure (b) constant volume
(c) constant temperature (d) constant temperature & pressure
- 3) In a closed thermos flask few cubes of ice are placed. It is example of.....
 (a) closed system (b) open system
(c) isolated system (d) thermodynamics system
- 4) In an adiabatic compression, the internal energy of the system.....
 (a) decreases **(b) increases** (c) increases or decreases (d) none
- 5) Identify the intensive quantities from the following.....
 (a) volume (b) enthalpy **(c) pressure** (d) none
- 6) Which of the following is extensive property?
 (a) mole (b) mass (c) volume **(d) all are correct**
- 7) The heat change in a chemical reaction at constant pressure is.....
(a) ΔH (b) ΔU (c) ΔT (d) ΔV
- 8) ΔU is always positive when.....
(a) system absorbs heat and work is done on it
 (b) system emits heat and work is done on it
 (c) system emits heat and no work is done on it
 (d) system absorbs heat and work is done by it
- 9) If in a chemical reaction, the products have less energy than the reactants, that reaction is.....
 (a) a reversible reaction **(b) an exothermic reaction**
 (c) an isothermal reaction (d) an endothermic reaction
- 10) A small piece of Na metal is dropped into water. Na bursts into flame, the reaction is:
 (a) neutralisation (b) reversible **(c) exothermic** (d) endothermic
- 11) For the reaction: $C_3H_8(g) + 5O_2(g) \rightarrow 3CO_2(g) + 4H_2O(l)$ at constant temperature, $\Delta H - \Delta U$ is
 (a) $+RT$ **(b) $-3RT$** (c) $+3RT$ (d) $-RT$
- 12) When NH_4Cl is dissolved in water, the solution becomes cold. The change is?
(a) endothermic (b) exothermic (c) super cooling (d) none

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The Editors of REA, Ralph Pike



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Thermodynamics DeMYSTiFied Merle C. Potter, 2009-03-03 Take the heat off of understanding thermodynamics Now

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Problems In Chemical Thermodynamics, With Solutions Maka Aleksishvili, Shota Sidamonidze, 2002-11-26 The methods of chemical thermodynamics are effectively used in many fields of science and technology Mastering these methods and their use in practice requires profound comprehension of the theoretical questions and acquisition of certain calculating skills This book is useful to undergraduate and graduate students in chemistry as well as chemical thermal and refrigerating technology it will also benefit specialists in all other fields who are interested in using these powerful methods in their practical activities

Basic Thermodynamics B. K. VENKANNA, SWATI B. V., B. K. Venkanna, B. V. Swati, 2010-07 This book provides an in depth discussion of the principles of thermodynamics It focuses on engineering applications of theory and sound techniques for solving thermodynamic problems The book presents the fundamental concepts of thermodynamics and describes the theory of work and heat The text covers in detail the first law and the second law of thermodynamics with their applications It also explains the concepts of entropy and availability and irreversibility In addition the book presents thermodynamic properties of pure substances ideal gases and mixtures of ideal gases as well as real gases This book is designed for undergraduate students of mechanical engineering industrial and production engineering automobile engineering and aeronautical engineering for their courses in thermodynamics Key Features Presents the text in a simple and elegant manner to enable the students to grasp the essentials of the subject easily and quickly Covers all types of problems of various difficulty levels Includes more than 300 worked out examples and a large number of end of chapter exercises Provides

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