Stoichiometry

Section 11.1 What is stoichiometry?

In your textbook, read about stoichiometry and the balanced equation.

For each statement below, write true or false.

- The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry.
 - 2. Stoichiometry is based on the law of conservation of mass.
 - In any chemical reaction, the mass of the products is less than the mass of the reactants.
- 4. The coefficients in a chemical equation represent not only the number of individual particles but also the number of moles of particles.
 - The mass of each reactant and product is related to its coefficient in the balanced chemical equation for the reaction by its molar mass.

Complete the table below, using information represented in the chemical equation for the combustion of methanol, an alcohol.

methanol + oxygen → carbon dioxide + water

 $2CH_3OH(1) + 3O_3(g) \rightarrow 2CO_3(g) + 4H_3O(g)$

Substance	Molar Mass (g/mol)	Number of Molecules	Number of Moles (mol)	Mass (g)
6. Methanol	32.05			
7. Oxygen gas	32.00		3	
8. Carbon dioxide	44.01		(i)	
9. Water	18.02			

- 10. What are the reactants?
- 11. What are the products?
- 12. What is the total mass of the reactants?
- 13. What is the total mass of the products?
- 14. How do the total masses of the reactants and products compare?

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Michael Levy, Barton A. Weitz, Dhruv Grewal Mcgraw Hill Retail Management Study Guide:

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Table of Contents Mcgraw Hill Retail Management Study Guide

- 1. Understanding the eBook Mcgraw Hill Retail Management Study Guide
 - The Rise of Digital Reading Mcgraw Hill Retail Management Study Guide
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Mcgraw Hill Retail Management Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcgraw Hill Retail Management Study Guide
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Mcgraw Hill Retail Management Study Guide
 - Personalized Recommendations
 - Mcgraw Hill Retail Management Study Guide User Reviews and Ratings
 - Mcgraw Hill Retail Management Study Guide and Bestseller Lists
- 5. Accessing Mcgraw Hill Retail Management Study Guide Free and Paid eBooks
 - Mcgraw Hill Retail Management Study Guide Public Domain eBooks
 - Mcgraw Hill Retail Management Study Guide eBook Subscription Services
 - Mcgraw Hill Retail Management Study Guide Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Mcgraw Hill Retail Management Study Guide Compatibility with Devices
 - Mcgraw Hill Retail Management Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mcgraw Hill Retail Management Study Guide
 - Highlighting and Note-Taking Mcgraw Hill Retail Management Study Guide
 - Interactive Elements Mcgraw Hill Retail Management Study Guide

- 8. Staying Engaged with Mcgraw Hill Retail Management Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcgraw Hill Retail Management Study Guide
- 9. Balancing eBooks and Physical Books Mcgraw Hill Retail Management Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcgraw Hill Retail Management Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mcgraw Hill Retail Management Study Guide
 - Setting Reading Goals Mcgraw Hill Retail Management Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mcgraw Hill Retail Management Study Guide
 - Fact-Checking eBook Content of Mcgraw Hill Retail Management Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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